

Chemistry Paper 1 revision

ELECTROLYSIS:

- *Separating elements in insoluble ionic compounds*

Equipment-

- *Beaker containing an electrolyte (liquid/solution containing an ionic compound) → ions are free to move*
- *Electrodes, which are solid conductors (usually metal or carbon), positive electrode (anode), negative electrode (cathode)*
- *Wire, used to join the electrodes so the electrons can flow between the electrodes*
- *Add a power supply to the wire, such as a battery to drive the flow of electrons*

Electrolysis- "splitting up with electricity", by using the battery to pass an electric current through the electrolyte.

Negatively charged ions go to the opposite charge such as the positively charged ions do, for example:

- *Negatively charged ions → Anode (positive electrolyte)*
- *Positively charged ions → Cathode (negative electrolyte)*

Once these ions go to the oppositely charged electrolyte it becomes discharged which means that it becomes a neutral atom such as Br⁻ would become Br (negatively charged bromine ion becomes bromine)

Electrons are reduced at a cathode, and are increased at an anode

PART TWO

Electrolysis to extract reactive metals from their oxides:

- Examples of metal oxides would be aluminium oxide and copper oxide, the goal is to separate the oxygen molecules from the metal, (remove the oxygen)
- One way this can be achieved is through reduction with carbon, where the carbon displaces the metal, and the equation would be metal oxide + carbon → pure metal + carbon dioxide. However, this only works for metals that are less reactive than carbon.

REACTIVITY SERIES-

Potassium

Sodium

Lithium

Calcium

Magnesium

CARBON

Zinc

Iron

Hydrogen

Copper

Chemistry Paper 1 06

AN Whitehead



Chemistry Paper 1 06:

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Table of Contents Chemistry Paper 1 06

1. Understanding the eBook Chemistry Paper 1 06
 - The Rise of Digital Reading Chemistry Paper 1 06
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Paper 1 06
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Paper 1 06
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Paper 1 06
 - Personalized Recommendations

- Chemistry Paper 1 06 User Reviews and Ratings
- Chemistry Paper 1 06 and Bestseller Lists
- 5. Accessing Chemistry Paper 1 06 Free and Paid eBooks
 - Chemistry Paper 1 06 Public Domain eBooks
 - Chemistry Paper 1 06 eBook Subscription Services
 - Chemistry Paper 1 06 Budget-Friendly Options
- 6. Navigating Chemistry Paper 1 06 eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Paper 1 06 Compatibility with Devices
 - Chemistry Paper 1 06 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Paper 1 06
 - Highlighting and Note-Taking Chemistry Paper 1 06
 - Interactive Elements Chemistry Paper 1 06
- 8. Staying Engaged with Chemistry Paper 1 06
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Paper 1 06
- 9. Balancing eBooks and Physical Books Chemistry Paper 1 06
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Paper 1 06
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Paper 1 06
 - Setting Reading Goals Chemistry Paper 1 06
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Paper 1 06
 - Fact-Checking eBook Content of Chemistry Paper 1 06

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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