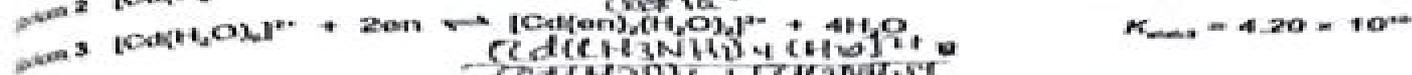
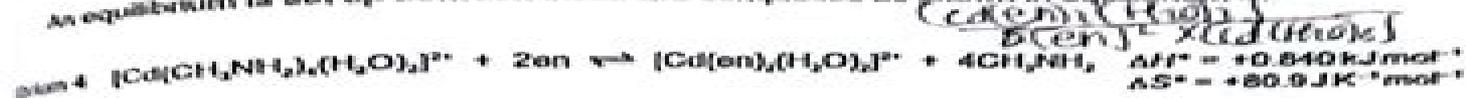


(b) Cadmium ions form complexes with methylamine, CH_3NH_2 , and with 1,2-diaminoethane, $\text{H}_2\text{NCH}_2\text{CH}_2\text{NH}_2$, as shown in equilibria 2 and 3. 1,2-diaminoethane is shown as en.



An equilibrium is set up between these two complexes as shown in equilibrium 4.



(i) K_{eq} is the equilibrium constant for equilibrium 4.

Write an expression for K_{eq} in terms of $K_{\text{stab}2}$ and $K_{\text{stab}3}$.

$$K_{\text{eq}} = \frac{[\text{CH}_3\text{NH}_2]^4 \times [\text{Cd}(\text{en})_2(\text{H}_2\text{O})_2]^{2+}}{[\text{en}]^2 \times [\text{Cd}(\text{CH}_3\text{NH}_2)_4(\text{H}_2\text{O})_2]^{2+}} = \frac{K_{\text{stab}3}}{K_{\text{stab}2}} \quad [1]$$

(ii) Calculate the value of the standard Gibbs free energy change, ΔG° , for equilibrium 4 at 298 K.

$$\Delta G = \Delta H - T\Delta S = +0.840 - 298 \times (80.9 \times 10^{-3})$$

$$\Delta G^\circ = -23.3 \text{ kJ mol}^{-1} \quad [2]$$

(iii) State how the value of ΔG° changes as the temperature increases. Explain your answer.

As temp increases, ΔG value gets more negative as value for $(T\Delta S)$ in the eqn increases. [1]

[Total: 9]

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