

# QUALITATIVE ANALYSIS

| Experiment   | Observation   | Inference  |
|--|---|--|
| <b><u>Preliminary Test</u></b>   |   |  |
| 1. Noted the colour of the given salt  | 1. Blue or Bluish Green<br>2. Greenish<br>3. Light Green<br>4. Dark Brown<br>5. Light Pink<br>6. Colourless   | $\text{Cu}^{2+}$ or $\text{Ni}^{2+}$<br>$\text{Ni}^{2+}$<br>$\text{Fe}^{2+}$<br>$\text{Fe}^{3+}$<br>$\text{Mn}^{2+}$<br>Absence of $\text{Cu}^{+2}$ , $\text{Ni}^{2+}$ , $\text{Fe}^{2+}$ , $\text{Fe}^{3+}$ , $\text{Mn}^{2+}$  |
| 2. Noted the solubility  | 1. Soluble in water<br>2. Soluble in dilute HCl   | $\text{NH}_4^+$ may be present<br>Absence of group I cations   |
| 3. Dry Heating Test:<br>Heated a pinch of the salt in a dry Test tube                                    | 1. Colour, odourless gas<br>2. Colourless gas with rotten egg smell<br>3. Colourless gas with vinegar smell<br>4. Colourless gas with ammonia smell<br>5. Reddish brown gas<br>6. Greenish yellow gas<br>7. Violet gas  | $\text{CO}_3^{2-}$ may be present<br>$\text{S}^{2-}$ may be present<br>$\text{CH}_3\text{COO}^-$ may be present<br>$\text{NH}_4^+$ may be present<br>$\text{NO}_3^-$ or $\text{Br}^-$ may be present<br>$\text{Cl}^-$ may be present<br>$\text{I}^-$ may be present  |
| <b><u>Identification of acid radical (Anion)</u></b>   |   |  |
| 1. Dilute Acid Test:<br>Treated a pinch of the substance With dilute $\text{H}_2\text{SO}_4$             | 1. Colourless, odourless gas with brisk effervescence which turned lime water milky.<br>2. Vinegar Smell<br>3. No characteristic Reaction   | $\text{CO}_3^{2-}$ present<br>$\text{CO}_3^{2-} + 2\text{H}^+ \rightarrow \text{CO}_2 + \text{H}_2\text{O}$<br>$\text{Ca}(\text{OH})_2 + \text{CO}_2 \rightarrow \text{CaCO}_3 + \text{H}_2\text{O}$<br>$\text{CH}_3\text{COO}^-$ Present<br>$\text{CH}_3\text{COO}^- + \text{H}^+ \rightarrow \text{CH}_3\text{COOH}$<br>Absence of $\text{CO}_3^{2-}$ , $\text{CH}_3\text{COO}^-$  |
| 2. Concentrated Acid Test:<br>Treated a pinch of the substance With concentrated $\text{H}_2\text{SO}_4$ | 1. Colourless gas with pungent Smell which gave dense white fumes when a glass rod dipped in $\text{NH}_4\text{OH}$ was brought near the mouth of the test tube<br>2. Reddish brown vapours with Pungent smell. Solution turns red<br>3. Deep violet vapours with pungent smell. Sides of test tube | $\text{Cl}^-$ present<br>$\text{Cl}^- + \text{H}^+ \rightarrow \text{HCl}$<br>$\text{HCl} + \text{NH}_4\text{OH} \rightarrow \text{NH}_4\text{Cl} + \text{H}_2\text{O}$<br>$\text{Br}^-$ present<br>$\text{Br}^- + \text{H}^+ \rightarrow \text{HBr}$<br>$2\text{HBr} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Br}_2 + \text{SO}_2$<br>$\text{I}^-$ present<br>$\text{I}^- + \text{H}^+ \rightarrow \text{HI}$<br>$2\text{HI} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{I}_2 + \text{SO}_2$ |

# Chemistry Practical Qualitative Analysis Sheet

**Amory Prescott Folwell**



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