

# Reference Tables for Physical Setting/CHEMISTRY

## 2002 Edition

**Table A**  
**Standard Temperature and Pressure**

Name	Value	Unit
Standard Pressure	101.3 kPa 1 atm	kilopascal atmosphere
Standard Temperature	273 K 0°C	kelvin degree Celsius

**Table B**  
**Physical Constants for Water**

Heat of Fusion	334 J/g
Heat of Vaporization	2260 J/g
Specific Heat Capacity of H <sub>2</sub> O (l)	4.18 J/g•°C

**Table C**  
**Selected Prefixes**

Factor	Prefix	Symbol
10 <sup>3</sup>	kilo-	k
10 <sup>-1</sup>	deci-	d
10 <sup>-2</sup>	centi-	c
10 <sup>-3</sup>	milli-	m
10 <sup>-6</sup>	micro-	μ
10 <sup>-9</sup>	nano-	n
10 <sup>-12</sup>	pico-	p

**Table D**  
**Selected Units**

Symbol	Name	Quantity
m	meter	length
g	gram	mass
Pa	pascal	pressure
K	kelvin	temperature
mol	mole	amount of substance
J	joule	energy, work, quantity of heat
s	second	time
L	liter	volume
ppm	part per million	concentration
M	molarity	solution concentration

**Table E**  
**Selected Polyatomic Ions**

H <sub>3</sub> O <sup>+</sup>	hydronium	CrO <sub>4</sub> <sup>2-</sup>	chromate
Hg <sub>2</sub> <sup>2+</sup>	mercurous (II)	Cr <sub>2</sub> O <sub>7</sub> <sup>2-</sup>	dichromate
NH <sub>4</sub> <sup>+</sup>	ammonium	MnO <sub>4</sub> <sup>-</sup>	permanganate
C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup> CH <sub>3</sub> COO <sup>-</sup>	acetate	NO <sub>2</sub> <sup>-</sup>	nitrite
CN <sup>-</sup>	cyanide	NO <sub>3</sub> <sup>-</sup>	nitrate
CO <sub>3</sub> <sup>2-</sup>	carbonate	O <sub>2</sub> <sup>2-</sup>	peroxide
HCO <sub>3</sub> <sup>-</sup>	hydrogen carbonate	OH <sup>-</sup>	hydroxide
C <sub>2</sub> O <sub>4</sub> <sup>2-</sup>	oxalate	PO <sub>4</sub> <sup>3-</sup>	phosphate
ClO <sup>-</sup>	hypochlorite	SCN <sup>-</sup>	thiocyanate
ClO <sub>2</sub> <sup>-</sup>	chlorite	SO <sub>3</sub> <sup>2-</sup>	sulfite
ClO <sub>3</sub> <sup>-</sup>	chlorate	SO <sub>4</sub> <sup>2-</sup>	sulfate
ClO <sub>4</sub> <sup>-</sup>	perchlorate	H <sub>2</sub> SO <sub>4</sub>	hydrogen sulfate
		S <sub>2</sub> O <sub>3</sub> <sup>2-</sup>	thiosulfate

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