

CHEMISTRY REVISION GUIDE

for IGCSE Coordinated Science

This revision guide is designed to help you study for the chemistry part of the IGCSE Coordinated Science course.

The guide contains everything that the syllabus says you need you need to know, and nothing extra.

The material that is only covered in the supplementary part of the course (which can be ignored by core candidates) is highlighted in dashed boxes:



Whilst this guide is intended to help with your revision, it should not be your only revision. It is intended as a starting point but *only a starting point*. You should make sure that you also read your text books and use the internet to supplement your study in conjunction with your syllabus document.

Whilst this guide does contain the entire syllabus, it just has the bare minimum and is not in itself sufficient for those candidates aiming for the highest grades. If that is you, you should make sure you read around a range of sources to get a deeper knowledge and understanding.

Some very useful websites to help you further your understanding include:

• <http://www.docbrown.info/> - whilst not the prettiest site this contains a lot of very useful and nicely explained information.

• <http://www.bbc.co.uk/schools/gcsebitesize/science/> - well presented with many clear diagrams, animations and quizzes. Can occasionally lack depth.

• <http://www.chemguide.co.uk/> - whilst mostly targeted at A-Levels this site contains very detailed information suitable for those looking to deepen their knowledge and hit the highest grades.

Finally, remember revision is not just reading but should be an active process and could involve:

- Making notes
- Condensing class notes
- Drawing Mind-maps
- Practicing past exam questions
- Making flashcards

The golden rule is that what makes you think makes you learn.

Happy studying, Mr Field.

C1: THE ELEMENTS OF CHEMISTRY

Atom: The smallest particle of matter	An atom: 	Some atoms:
Molecule: A small particle made from more than one atom bonded together	Molecules of an element: 	Molecules of a compound:
Element: A substance made of only one type of atom	A solid element: 	A gaseous element:
Compound: A substance made from two or more different elements bonded together	A solid compound: 	A gaseous compound:
Mixture: A substance made from two or more elements or compounds mixed but not joined	A mixture of compounds and elements: 	

CHEMICAL FORMULAS

Formulas tell you the atoms that make up a compound

Eg 1. H_2O - two H, one O
Eg 2. C_2H_6O - two C, six H, one O

Eg 3. $Mg(OH)_2$ - one Mg, two O, two H

Eg 4. $CH_2(CH_2)_2$ - three C, 8 H

* In this case everything in brackets is doubled

CHEMICAL EQUATIONS

-Show the reactants you start with and the products you make

-Must contain an arrow (\rightarrow) NOT an equals sign (=)

-Must be balanced - same number of atoms on each side



This is unbalanced as there are 4 'H' on the left but only 2 'H' on the right. This must be corrected by adding a '2' in front of the ' H_2O ':



Now the H balances but there 4 O on the right and only 2 on the left. This must be balanced by placing a '2' in front of the ' O_2 '



Now there is 1 'C', 4 'H' and 4 'O' on each side so it balances.

MOLES AND MOLAR MASS

-A 'mole' is the same we give to the number

6.02×10^{23} - it is used to talk about the numbers of particles involved in chemical reactions.

-It is the number of atoms such that one mole has the same mass in grams as an atom's atomic mass.

-Eg 1: Carbon has an atomic mass of 12.0, so a mole of carbon has a mass of 12.0g

-Eg 2: Iron has an atomic mass of 55.8, so a mole of iron has a mass of 55.8g

-The molar mass of a compound is the sum of the

masses of the elements in its formula:

-Eg 1: C_2H_6O (C=12.0, H=1.0, O=16.0)

Molar mass = $12.0 \times 2 + 1.0 \times 6 + 16.0 \times 1 = 46.0g$

The number of moles of a substance present in a given mass is given by:

$$\text{Moles} = \frac{\text{mass used}}{\text{molar mass}}$$

Eg: How many moles of ethanol (C_2H_6O) are present in 69.0g?

$$\text{Moles} = \frac{\text{mass used}}{\text{molar mass}} = \frac{69.0}{46.0} = 1.5 \text{ mol}$$

ATOMIC STRUCTURE

Atoms are made of:

Protons: mass = 1, charge = +1

Neutrons: mass = 1, charge = 0

Electrons: mass = 0, charge = -1

is a square on the periodic table the smaller number, the proton number gives the number of protons or electrons and the bigger number, the nucleon number the number of protons and neutrons together.



Eg 1: Boron has 5 protons, 6 neutrons, 5 electrons

Eg 2: Phosphorus has 15 protons, 16 neutrons, 15 electrons

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