

## Introduction 11

# An Electron's Address

Name: \_\_\_\_\_  
 Date: \_\_\_\_\_  
 Class: \_\_\_\_\_

### Introduction: Energy Levels and Sublevels

In previous lessons you have learned about the Bohr model of the atom. The Bohr model of the atom says that electrons orbit the nucleus in fixed energy levels. The Bohr model of the atom says that electrons orbit the nucleus in fixed energy levels. Bohr's model of the atom says that electrons orbit the nucleus in fixed energy levels.

Bohr's model of the atom says that electrons orbit the nucleus in fixed energy levels. Bohr's model of the atom says that electrons orbit the nucleus in fixed energy levels. Bohr's model of the atom says that electrons orbit the nucleus in fixed energy levels.

Table 1

Energy Level	Number of Sublevels in Each Energy Level
1	1
2	2
3	3
4	4

How does the number of sublevels in each energy level change as the energy level increases? Does the number of sublevels in each energy level change as the energy level increases?

### Critical Thinking Questions

- How many sublevels are in the 1<sup>st</sup> energy level?  
 The number of sublevels in the 1<sup>st</sup> energy level is 1. The number of sublevels in the 1<sup>st</sup> energy level is 1.
- How many sublevels are in the 2<sup>nd</sup> energy level?  
 The number of sublevels in the 2<sup>nd</sup> energy level is 2. The number of sublevels in the 2<sup>nd</sup> energy level is 2.
- How many sublevels are in the 3<sup>rd</sup> energy level?  
 The number of sublevels in the 3<sup>rd</sup> energy level is 3. The number of sublevels in the 3<sup>rd</sup> energy level is 3.
- How many sublevels are in the 4<sup>th</sup> energy level?  
 The number of sublevels in the 4<sup>th</sup> energy level is 4. The number of sublevels in the 4<sup>th</sup> energy level is 4.
- How many sublevels are in the 5<sup>th</sup> energy level?  
 The number of sublevels in the 5<sup>th</sup> energy level is 5. The number of sublevels in the 5<sup>th</sup> energy level is 5.

# Chemquest 11 Answers

**Siddappa N.Byrareddy**



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