

17. 1.69×10^{22} carbon atoms; 9.87×10^{23} hydrogen atoms;
 4.25×10^{21} chlorine atoms; 2.82×10^{21} oxygen atoms

18. a. 111.11 g/mol

b. 1.20×10^{25} carbon atoms; 1.20×10^{25} hydrogen atoms; 4.82×10^{24} oxygen atoms; 2.41×10^{24} nitrogen atoms

19. The order by mass is: carbon, oxygen, nitrogen, hydrogen. The order by number of atoms is: hydrogen, carbon, oxygen, nitrogen. The orders are not the same, because atoms of different elements have different masses.

20. To convert number of particles of a substance to amount of moles, divide by the Avogadro constant. To convert amount of moles of a substance to number of particles, multiply by the Avogadro constant.

21. a. 4.25 kg b. 14.5 kg

22. They would all have the same number of molecules, because each substance has the same number of molecules per mole (the Avogadro constant).

23. 57.51 g ethanol, 22.49 g water

24. a. Measure the mass of the one dozen marbles and divide by 12. Note: Students may, for simplicity, assume the mass of the plastic bag is negligible, or add a step to weigh a sample of plastic bags and determine an average mass to be subtracted from the mass of the sample.

b. First method—Multiply the mass of one marble by 365.

Second method—Divide 365 by 12 to determine how many dozen marbles are in the jar. Multiply the number by the mass of one dozen marbles.

c. First method—Multiply the mass of one marble by 6.02×10^{23} .

Second method—Divide 6.02×10^{23} by 12 to find how many dozen marbles are in one mole. Then, multiply the number of dozen marbles by the mass of one dozen to get the mass of one mole of marbles.

Measuring the mass of one dozen marbles is like calculating the molecular molar mass of a substance with 12 atoms per molecule. To get the mass of one molecule, you would divide the molar mass by the Avogadro constant, just like you divide the mass of the dozen marbles by 12 to get the mass of one marble.

To find the specific mass of a particular number of particles of a substance, you would divide the number of particles by the Avogadro constant to get the amount of moles, and then multiply by the molar mass to determine the mass.

25. 119 pg/day

Chapter 6 Proportions in Chemical Compounds

Learning Check Questions

(Student textbook page 262)

1. Yes, water is the same regardless of the source. We know this because of the law of definite proportions.
2. Elements will form bonds with other elements until all of their valence shells are full. This can be achieved by many proportions of the same elements.
3. Both sets of terms (definite and constant, and proportion and composition) indicate that the ratio of elements in a compound remains the same.
4. No, the mass percent of carbon in carbon dioxide cannot change, based on the law of definite composition.
5. Yes, because the carbon is present in different ratios in each compound.
6. a. 57.1%
b. Carbon and oxygen have different mass percents because their molar masses are different.

(Student textbook page 270)

7. There is one atom of carbon for every four atoms of hydrogen, so the ratio is 1:4.
8. Both provide information about the proportions of elements found in the compound. However, the proportions in a percentage composition are based on the overall mass of each element found in the compound, while the proportions in a molecular formula are based on the number of atoms of each element in the compound.
9. The molecular formula shows the actual number of atoms of each element in a molecule of the substance, while the empirical formula only tells you the ratios of the atoms in a molecule.
10. The molecular and empirical formulas are the same when the actual number of atoms of each element in a compound is already in the lowest whole number ratio.

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Table of Contents Mhr Chemistry 11 Solutions

1. Understanding the eBook Mhr Chemistry 11 Solutions
 - The Rise of Digital Reading Mhr Chemistry 11 Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Mhr Chemistry 11 Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mhr Chemistry 11 Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mhr Chemistry 11 Solutions
 - Personalized Recommendations
 - Mhr Chemistry 11 Solutions User Reviews and Ratings
 - Mhr Chemistry 11 Solutions and Bestseller Lists

-
5. Accessing Mhr Chemistry 11 Solutions Free and Paid eBooks
 - Mhr Chemistry 11 Solutions Public Domain eBooks
 - Mhr Chemistry 11 Solutions eBook Subscription Services
 - Mhr Chemistry 11 Solutions Budget-Friendly Options
 6. Navigating Mhr Chemistry 11 Solutions eBook Formats
 - ePub, PDF, MOBI, and More
 - Mhr Chemistry 11 Solutions Compatibility with Devices
 - Mhr Chemistry 11 Solutions Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mhr Chemistry 11 Solutions
 - Highlighting and Note-Taking Mhr Chemistry 11 Solutions
 - Interactive Elements Mhr Chemistry 11 Solutions
 8. Staying Engaged with Mhr Chemistry 11 Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mhr Chemistry 11 Solutions
 9. Balancing eBooks and Physical Books Mhr Chemistry 11 Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mhr Chemistry 11 Solutions
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Mhr Chemistry 11 Solutions
 - Setting Reading Goals Mhr Chemistry 11 Solutions
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Mhr Chemistry 11 Solutions
 - Fact-Checking eBook Content of Mhr Chemistry 11 Solutions
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning

-
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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