

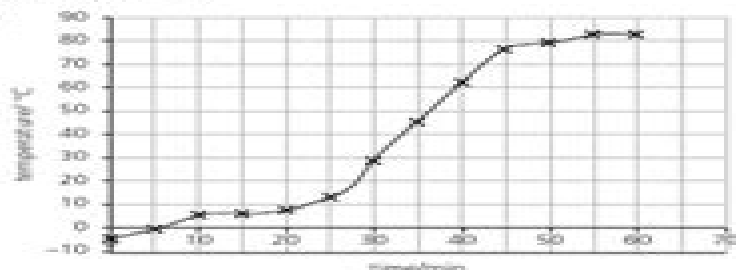
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Answers

1 All about matter

- 1 (a) **A** condensation, **B** boiling, **C** freezing, **D** melting, **E** sublimation [1 for each]
 (b) gas – particles are far apart and move around randomly [1], colliding with each other occasionally [1]
 liquid – particles are closer together [1], they move relative to one another [1], colliding more often than in the case of gases [1]
 solid – particles are very close together [1] and vibrate about a fixed point [1]
 (c) (i) carbon dioxide or iodine [1]
 (ii) water [1]

2 (a)



[1 for scale, 1 for axis labels and units, 1 for points plotted correctly, 1 for line]

(b) 5 °C [1]

(c) 82 °C [1]

(d) The particles move around more as they gain kinetic energy. This breaks more of the inter-particle bonds, allowing the particles to move away from each other. [2]

- 3 In ice, I (the water molecule) am bonded to other water molecules and vibrating in a fixed position [1]. As the temperature rises, there is enough energy for some of the bonds between us to weaken and begin to break during the increase in vibrations [1]. This allows me to move further from my neighbouring molecules and I am able to move past them [1]. As the temperature continues to rise, more of the bonds between us begin to break and I move away from my neighbouring molecules and move much more quickly [1]. At 100 °C I have so little attraction to my neighbouring molecules that I am able to break away [1] and move a long way away from them [1].
- 4 (a) converting temperatures to 295 K and 299 K [1]
 $V = 38 \times 299/295$ [1] = 38.51 cm³ [1]
 (b) converting temperatures to 293 K and 303 K [1]
 $V = 75 \times 303/293$ [1] = 77.56 cm³ [1]
 (c) temperatures are 298 K and 350 K [1]
 $V = (100 \times 350)/(2 \times 298)$ [1] = 58.72 cm³ [1]

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