

Campbell Biology, 11e (Urry)
Chapter 3 Water and Life

3.1 Multiple-Choice Questions

1) In a single molecule of water, two hydrogen atoms are bonded to a single oxygen atom by _____.

- A) hydrogen bonds
- B) nonpolar covalent bonds
- C) polar covalent bonds
- D) ionic bonds

Answer: C

Bloom's Taxonomy: Knowledge/Comprehension

Section: 3.1

2) The partial negative charge at one end of a water molecule is attracted to the partial positive charge of another water molecule. What is this attraction called?

- A) a covalent bond
- B) a hydrogen bond
- C) an ionic bond
- D) a van der Waals interaction

Answer: B

Bloom's Taxonomy: Knowledge/Comprehension

Section: 3.1

3) The partial negative charge in a molecule of water occurs because _____.

- A) the oxygen atom donates an electron to each of the hydrogen atoms
- B) the electrons shared between the oxygen and hydrogen atoms spend more time around the oxygen atom nucleus than around the hydrogen atom nucleus
- C) the oxygen atom has two pairs of electrons in its valence shell that are not neutralized by hydrogen atoms
- D) one of the hydrogen atoms donates an electron to the oxygen atom

Answer: B

Bloom's Taxonomy: Knowledge/Comprehension

Section: 3.1

4) Sulfur is in the same column of the periodic table as oxygen, but has electronegativity similar to carbon. Compared to water molecules, molecules of H₂S will _____.

- A) have greater cohesion to other molecules of H₂S
- B) have a greater tendency to form hydrogen bonds with each other
- C) have a higher capacity to absorb heat for the same change in temperature
- D) not form hydrogen bonds with each other

Answer: D

Bloom's Taxonomy: Synthesis/Evaluation

Section: 3.1

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