



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either an ionic or a molecular compound.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - C_2H_6
 - H_2SO_4
 - CaCl_2
 - $\text{C}_2\text{H}_5\text{OH}$
 - CO_2
 - H_2O_2
 - CH_4
 - H_2S

- Classify each of the following as either an ionic or a molecular compound.
 - K_2SO_4
 - C_2H_6
 - H_2O
 - HCl
 - CaCl_2
 - $\text{C}_2\text{H}_5\text{OH}$
 - CO_2
 - H_2O_2
 - CH_4
 - H_2S
- Write the formula for each of the following molecular compounds.
 - Carbon dioxide
 - Hydrogen peroxide
 - Hydrogen cyanide
 - Carbon monoxide
 - Diphosphorus pentoxide
 - Diphosphorus trioxide
 - Diphosphorus pentasulfide
 - Diphosphorus trisulfide
 - Diphosphorus tetrasulfide
 - Diphosphorus hexasulfide
 - Diphosphorus heptasulfide
 - Diphosphorus octasulfide
 - Diphosphorus nonasulfide
 - Diphosphorus decasulfide
 - Diphosphorus undecasulfide
 - Diphosphorus dodecasulfide
 - Diphosphorus trisulfide
 - Diphosphorus tetrasulfide
 - Diphosphorus pentasulfide
 - Diphosphorus hexasulfide
 - Diphosphorus heptasulfide
 - Diphosphorus octasulfide
 - Diphosphorus nonasulfide
 - Diphosphorus decasulfide
 - Diphosphorus undecasulfide
 - Diphosphorus dodecasulfide

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride. $\text{H}-\text{F}$
- Show covalent bonding in methane. C with four H atoms around it.
- Show covalent bonding in carbon dioxide. $\text{O}=\text{C}=\text{O}$
- Show covalent bonding in acetylene. $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for each of the following.
 - H_2O
 - H_2O_2
 - CO_2
 - C_2H_6
 - C_2H_4
 - C_2H_2
 - $\text{C}_2\text{H}_6\text{O}$
 - $\text{C}_2\text{H}_6\text{O}_2$
 - $\text{C}_2\text{H}_4\text{O}_2$
 - $\text{C}_2\text{H}_2\text{O}_2$
 - $\text{C}_2\text{H}_4\text{O}$
 - $\text{C}_2\text{H}_2\text{O}$
 - $\text{C}_2\text{H}_6\text{N}_2$
 - $\text{C}_2\text{H}_4\text{N}_2$
 - $\text{C}_2\text{H}_2\text{N}_2$
 - $\text{C}_2\text{H}_6\text{N}_2\text{O}$
 - $\text{C}_2\text{H}_4\text{N}_2\text{O}$
 - $\text{C}_2\text{H}_2\text{N}_2\text{O}$
 - $\text{C}_2\text{H}_6\text{N}_2\text{O}_2$
 - $\text{C}_2\text{H}_4\text{N}_2\text{O}_2$
 - $\text{C}_2\text{H}_2\text{N}_2\text{O}_2$

SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and suborbital are represented by each of the following?
 - $2s$
 - $2p_x$
 - $2p_y$
 - $2p_z$
 - $3s$
 - $3p_x$
 - $3p_y$
 - $3p_z$
 - $4s$
 - $4p_x$
 - $4p_y$
 - $4p_z$
 - $5s$
 - $5p_x$
 - $5p_y$
 - $5p_z$
- Which orbital and suborbital are represented by each of the following?
 - $3d_{xy}$
 - $3d_{yz}$
 - $3d_{zx}$
 - $3d_{x^2-y^2}$
 - $3d_{z^2}$
 - $4d_{xy}$
 - $4d_{yz}$
 - $4d_{zx}$
 - $4d_{x^2-y^2}$
 - $4d_{z^2}$
 - $5d_{xy}$
 - $5d_{yz}$
 - $5d_{zx}$
 - $5d_{x^2-y^2}$
 - $5d_{z^2}$

Chapter 8 Covalent Bonding Practice Problems Answers

Fred M. Dewey



Chapter 8 Covalent Bonding Practice Problems Answers:

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the GRE Chemistry Subject Test Princeton Review (Firm),2005 The GRE subject tests are among the most difficult standardized exams Rather than testing general problem solving skills they require highly specialized knowledge The experts at The Princeton Review have thoroughly research each subject test to provide students with the most thorough up to date information available Students don t need to relearn the entire histories of their fields just what they need to know to earn high scores on the exams Each guide includes one full length practice exam complete with comprehensive explanations for every solution **Basic Concepts of Chemistry** Alan Sherman,Sharon Sherman,Leonard Russikoff,1984 Basic

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