

MODULE C: SOLIDS, LIQUIDS & GASES

Kinetic Molecular Theory & Phases of Matter

- The different phases of matter can be described based on the Kinetic Molecular Theory

a) According to Kinetic theory, an increase in temperature will cause the average kinetic energy of the particles in the solid, liquid, and gas to increase.

↑ temperature ↑ average kinetic energy

b) When a solid is heated, the temperature is raised and the solid changes to a liquid at its melting point. If the heating is continued, the liquid will change into a gas at its boiling point.

c) The heat energy is changed into the equivalent amount of kinetic energy. Thus, the kinetic energy content of the particles increases in the order:

solid < liquid < gas

- Kinetic Molecular Theory

- 1) Matter is made up of particles
- 2) The particles are always moving

- Kinetic energy

→ energy of movement

- Kinetic Molecular Theory for Gases

a) A gas consists of small particles (atoms or molecules) which have negligible volume (V of gas molecule ≈ 0). This implies that:

$$V_{\text{container}} = V_{\text{molecules}} + V_{\text{free space}} \approx V_{\text{free space}}$$

and that there is plenty of space for compression.

b) Gas particles are separated relatively far from one another.

c) Gas particles behave independently from one another. Thus, inter-particle forces of attraction are negligible.

d) The gas particles are in rapid and continuous motion.

e) In the gaseous state, the gas particles have vibrational, rotational, and translational motion.

1) Vibrational motion - a particle vibrates about a fixed point

2) Rotational motion - a particle spins about a fixed point

3) Translational motion - a particle moves from one place to another

f) Pressure of a gas comes from molecules colliding with the walls of the container.

Such collisions are perfectly elastic.

g) Average kinetic energy depends only on the absolute temperature in Kelvin (K)

$$\text{Average Kinetic Energy} = \frac{3}{2} RT$$

R → gas law constant ($0.08206 \frac{\text{L-atm}}{\text{K-mol}}$)
 T → temperature (in Kelvin)

Chapter Review Part B Concept Answers

Dennis Zill, Jacqueline Dewar



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Glencoe Science McGraw-Hill Staff, 2001-08

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Appendix D show examples similar to the EHRs you will encounter in the workplace NEW Updated content includes the latest coding information available promoting accurate coding and success on the job NEW Additional exercise questions covering the Official Guidelines for Coding and Reporting Glencoe Sci Earth Science Chapter 21 Our Impact on Water and Air Chp Res 519 02 McGraw-Hill Staff,2001-08 Algebra and Trigonometry Dennis Zill,Jacqueline Dewar,2010-12 Written for a one or two term course at the freshman sophomore level the third edition covers the principles of college algebra trigonometry and analytic geometry in the concise and student friendly style that have made Zill s texts a world wide success It includes all of the trademark features for which Zill is known including lucid examples and problem sets a rich pedagogy a complete teaching and learning ancillary package and much more Throughout the text readers will find a wide range of word problems and relevant applications historical accounts of famous mathematicians and a strong variety of modern exercises

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aug 3 2023 minor chords are the set of chords that contain a characteristic minor third interval away from the root they re typically associated with sad or melancholic feelings the simplest form of a minor chord is the minor triad which consists of the root minor third and perfect fifth intervals here s the basic minor triad in a minor

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a minor triad has a minor third m3 on the bottom a major third m3 on top and a perfect fifth p5 between the outer notes in harmonic analysis and on lead sheets a c minor chord can be notated as cm c cmin or simply the lowercase c a minor triad is represented by the integer notation 0 3 7 a minor triad can also be described by its intervals the interval between the

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the a and c sections d and f on the lead sheet are similar to rhythm changes but with minor instead of major tonic chords on the first and third measures on the bridge the changes are similar to the b section of the head simplified in the first