

Chapter 10 test**Matching**

Match each item with the correct statement below.

- | | |
|----------------------------|--------------------------------------|
| a. representative particle | d. percent composition |
| b. mole | e. standard temperature and pressure |
| c. Avogadro's number | f. empirical formula |

- _____ 1. the number of representative particles of a substance present in 1 mole of that substance
- _____ 2. an atom, an ion, or a molecule, depending upon the way a substance commonly exists
- _____ 3. the SI unit used to measure amount of substance
- _____ 4. 0°C and 1 atm
- _____ 5. the percent by mass of each element in a compound
- _____ 6. the smallest whole number ratio of the atoms in a compound

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- _____ 7. What SI unit is used to measure the number of representative particles in a substance?
- | | |
|-------------|-----------|
| a. kilogram | c. kelvin |
| b. ampere | d. mole |
- _____ 8. How many hydrogen atoms are in 5 molecules of isopropyl alcohol, C₃H₇O?
- | | |
|-------------------------------------|--------------------------------------|
| a. $5 \times (6.02 \times 10^{23})$ | c. 35 |
| b. 5 | d. $35 \times (6.02 \times 10^{23})$ |
- _____ 9. All of the following are equal to Avogadro's number EXCEPT _____
- | |
|--|
| a. the number of atoms of bromine in 1 mol Br ₂ |
| b. the number of atoms of gold in 1 mol Au |
| c. the number of molecules of nitrogen in 1 mol N ₂ |
| d. the number of molecules of carbon monoxide in 1 mol CO |
- _____ 10. How many moles of tungsten atoms are in 4.8×10^{23} atoms of tungsten?
- | | |
|-------------------------------|-------------------------------|
| a. 8.0×10^{-2} moles | c. 1.3×10^{-4} moles |
| b. 8.0×10^1 moles | d. 1.3×10^{-2} moles |
- _____ 11. How many atoms are in 0.075 mol of titanium?
- | | |
|-------------------------|-------------------------|
| a. 1.2×10^{23} | c. 6.4×10^2 |
| b. 2.2×10^{24} | d. 4.5×10^{22} |
- _____ 12. How many molecules are in 2.10 mol CO₂?
- | | |
|------------------------------------|-------------------------------------|
| a. 2.53×10^{24} molecules | c. 3.49×10^{-24} molecules |
| b. 3.79×10^{24} molecules | d. 1.26×10^{24} molecules |

Chemistry Chapter 10 Test Answers

Hussin A.Rothana



Chemistry Chapter 10 Test Answers:

CliffsNotes HESI A2 Science Cram Plan Michael Reid,2021-04-13 A study guide for the HESI A2 science nursing school test that calendarizes a study plan for test takers depending on how much time they have left before taking the test Get a plan and make the most of the time you have left Whether you have two months one month or one week left before the exam you can turn to the experts at CliffsNotes for a trusted and achievable cram plan to ace the HESI A2 Science without ever breaking a sweat First you ll determine exactly how much time you have left to prepare for the exam Then you ll turn to the two month one month or one week cram plan for week by week and day by day schedules of the best way to focus your study according to your unique timeline Each stand alone plan includes Diagnostic tests help you pinpoint your strengths and weaknesses so you can focus your review on the topics in which you need the most help Subject areas review of material you should know for the exam biology chemistry anatomy and physiology and physics Practice exams with answers and detailed explanations McGraw-Hill's SAT Subject Test Chemistry, 3rd Edition Thomas Evangelist,2011-02-03 Expert guidance on the Chemistry exam Many colleges and universities require you to take one or more SAT II Subject Tests to demonstrate your mastery of specific high school subjects McGraw Hill s SAT Subject Test Chemistry is written by experts in the field and gives you the guidance you need perform at your best This book includes 4 full length sample tests updated for the latest test format 40 top tips to remember on test day Glossary of tested chemistry terms and formulas Tips and strategies from one of the most popular teachers at the renowned Brooklyn Technical High School Diagnostic test to pinpoint strengths and weaknesses Step by step review of all topics covered on the exam In depth coverage of the lab experiment questions that are a major test feature Charts tables and illustrations to simplify and reinforce learning Practice tests just like the real SAT Subject Test in Chemistry Test taking tips and strategies McGraw-Hill's SAT Subject Test: Chemistry, 2ed Thomas Evangelist,2009-02-01 We want to help you score high on the SAT Chemistry test We ve put all of our proven expertise into McGraw Hill s SAT Subject Test Chemistry to make sure you re fully prepared for this difficult exam With this book you ll get essential skill building techniques and strategies created by a leading high school chemistry teacher You ll also get 5 full length practice tests hundreds of sample questions and all the facts about the current exam With McGraw Hill s SAT Subject Test Chemistry we ll guide you step by step through your preparation program and give you the tools you need to succeed 5 full length sample tests with complete explanations for every question 40 top test items to remember on exam day A step by step review of all topics covered on the exam Teacher recommended tips and strategies to help you raise your score

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