

Test "Chemistry"

Chapter # 3

Objective part

Time: 15 min

Total marks: 12

Q.1- Every question has four possible answers. Encircle the correct option.

- i) **The atomic radii of the elements in periodic table;**
a) increase from left to right in a period
b) increase from top to bottom in a group
c) do not change from left to right in a period
d) decrease from top to bottom in a group
- ii) **Mendeleev Periodic table was based upon the:**
a) electronic configuration
b) atomic mass
c) atomic number
d) electro negativity
- iii) **The amount of energy given out when an electron is added to an atom is called;**
a) ionization energy
b) lattice energy
c) electron affinity
d) electro negativity
- iv) **Long form of periodic table is constructed on the basis of:**
a) Mendeleev postulate
b) atomic number
c) atomic mass
d) mass number
- v) **Along the period which of the following decreases:**
a) electro negativity
b) ionization energy
c) atomic radius
d) electron affinity
- vi) **Mark the incorrect statement about the ionization energy:**
a) it is measured kJmol^{-1}
b) it is the absorption of energy
c) it decreases in a period
d) it decreases in groups
- vii) **Transition elements are:**
a) all gases
b) all metals
c) all non metals
d) all metalloids
- viii) **S-f series of inner transition elements are called:**
a) lanthanides
b) actinides
c) halogens
d) alkali metals
- ix) **If we move from left to right in a period the value of ionization energy:**
a) remain same
b) decreases
c) increases
d) not effected
- x) **Which scientist has given the idea of actives in periodic table:**
a) Mendeleev
b) Al-razi
c) Newland
d) dobernier
- xi) **Shielding effect is greater in atoms with greater number of:**
a) proton
b) electron
c) neutron
d) positron
- xii) **Which is the incomplete period in the periodic table:**
a) 4th period
b) 2nd period
c) 7th period
d) 1st period

Chemistry Chapter 3 Test Answers

Lawrence Solomon



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