

Chapter 8

Organic Chemistry – Some Basic Principles and Techniques

Question 8.1:

What are hybridisation states of each carbon atom in the following compounds?

$\text{CH}_2 = \text{C} = \text{O}$, $\text{CH}_3\text{CH} = \text{CH}_2$, $(\text{CH}_3)_2\text{CO}$, $\text{CH}_2 = \text{CHCN}$, C_6H_6

Answer 8.1:

- (i) $\overset{1}{\text{C}}\text{H}_2 = \overset{2}{\text{C}} = \text{O}$
C-1 is sp^2 hybridised.
C-2 is sp hybridised.
- (ii) $\overset{1}{\text{C}}\text{H}_3 - \overset{2}{\text{C}}\text{H} = \overset{3}{\text{C}}\text{H}_2$
C-1 is sp^3 hybridised.
C-2 is sp^2 hybridised.
C-3 is sp^2 hybridised.
- (iii) $\overset{1}{\text{C}}\text{H}_3 - \overset{2}{\text{C}} = \overset{\text{O}}{\parallel} - \overset{3}{\text{C}}\text{H}_3$
C-1 is sp^3 hybridised.
- (iv) $\overset{1}{\text{C}}\text{H}_2 = \overset{2}{\text{C}}\text{H} - \overset{3}{\text{C}} = \text{N}$
C-2 is sp^2 hybridised.
C-3 is sp hybridised.
- (v) C_6H_6
All the 6 carbon atoms in benzene are sp^2 hybridised.

Question 8.2:

Indicate the σ and π bonds in the following molecules :

C_6H_6

C_6H_{12}

CH_2Cl_2

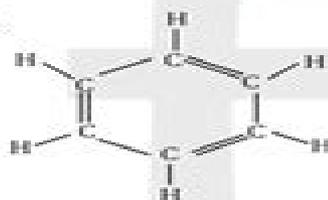
$\text{CH}_2 = \text{C} = \text{CH}_2$

CH_3NO_2

HCONHCH_3

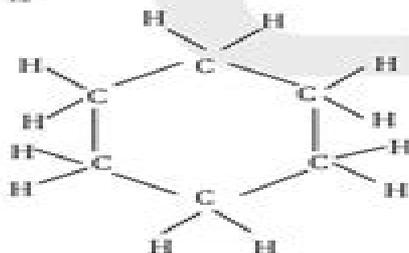
Answer 8.2:

(i) C_6H_6



There are six C-C sigma (σ_{C-C}) bonds, six C-H sigma (σ_{C-H}) bonds and three C=C pi (π_{C-C}) resonating bonds in this compound.

(ii) C_6H_{12}



There are six C-C sigma (σ_{C-C}) bonds and twelve C-H sigma (σ_{C-H}) bonds in this compound.

Chemistry Chapter 8

ML Morrison



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