

Stoichiometry Worksheet # 1

1) Glucose is used as a source of energy by the human body. The overall reaction in the body is



Calculate the number of moles of oxygen needed to oxidize 12.5 g of $\text{C}_6\text{H}_{12}\text{O}_6$.

2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH_3) can be produced?

3) Assume that in the decomposition of potassium chlorate, KClO_3 ,



80.5 moles of O_2 form. How many grams of potassium chlorite, the other product would be formed?

4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

5) The compound "Cisplatin" $\text{PtCl}_2(\text{NH}_3)_2$, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g K_2PtCl_4 ?

B. How many moles of NH_3 would be needed if 3.8 moles of K_2PtCl_4 were present?

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Table of Contents Pearson Education Chapter 12 Stoichiometry Answer Key

1. Understanding the eBook Pearson Education Chapter 12 Stoichiometry Answer Key
 - The Rise of Digital Reading Pearson Education Chapter 12 Stoichiometry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Education Chapter 12 Stoichiometry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Education Chapter 12 Stoichiometry Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Education Chapter 12 Stoichiometry Answer Key
 - Personalized Recommendations
 - Pearson Education Chapter 12 Stoichiometry Answer Key User Reviews and Ratings
 - Pearson Education Chapter 12 Stoichiometry Answer Key and Bestseller Lists

5. Accessing Pearson Education Chapter 12 Stoichiometry Answer Key Free and Paid eBooks
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 - Pearson Education Chapter 12 Stoichiometry Answer Key Budget-Friendly Options
6. Navigating Pearson Education Chapter 12 Stoichiometry Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Pearson Education Chapter 12 Stoichiometry Answer Key Compatibility with Devices
 - Pearson Education Chapter 12 Stoichiometry Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Education Chapter 12 Stoichiometry Answer Key
 - Highlighting and Note-Taking Pearson Education Chapter 12 Stoichiometry Answer Key
 - Interactive Elements Pearson Education Chapter 12 Stoichiometry Answer Key
8. Staying Engaged with Pearson Education Chapter 12 Stoichiometry Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Education Chapter 12 Stoichiometry Answer Key
9. Balancing eBooks and Physical Books Pearson Education Chapter 12 Stoichiometry Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Education Chapter 12 Stoichiometry Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pearson Education Chapter 12 Stoichiometry Answer Key
 - Setting Reading Goals Pearson Education Chapter 12 Stoichiometry Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson Education Chapter 12 Stoichiometry Answer Key
 - Fact-Checking eBook Content of Pearson Education Chapter 12 Stoichiometry Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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