

Prentice Hall Physical Science Chapter 4 Atomic Structure

4.1 Studying Atoms

- A. Ancient Greek Models of Atoms
 - Democritus believed all matter was made of extremely small particles that could not be divided
 - he called the particles atoms (from the Greek atomos meaning indivisible)
 - he thought there were different types of atoms with different properties (liquids had round, smooth atoms and solids were rough and prickly)
- B. Dalton's Atomic Theory
 - he proved atoms exist by discovering that compounds have a fixed composition
 - Dalton's Theory:
 - all elements are composed of atoms
 - all atoms of the same element have the same mass, and atoms of different elements have different masses
 - compounds contain atoms of more than one element
 - in a particular compound, atoms of different element always combine in the same way
 - he pictured atoms as solid spheres
 - in time it was found that not all of this was completely correct
- C. Thomson's Model of the Atom
 - some objects have a positive or negative charge and objects with the same charge repel and objects with opposite charges attract
 - Thomson used an electric current to learn more about atoms
 - he passed an electric current through a gas and a glowing beam appeared
 - the beam was repelled by a negatively charged metal plate (like charges repel and opposite charges attract)
 - Thomson concluded that the beam must be made of negatively charged particles, smaller than atoms.
 - since the atom was known to be electrically neutral, he developed a model in which negative charges were evenly scattered throughout an atom filled with a positively charged pudding like material
 - called the "Plum Pudding" model (or chocolate chip cookie dough)
- D. Rutherford's Atomic Theory
 - Rutherford discovered positively charged, fast moving particles called alpha particles
 - he shot some of these particles a piece of gold foil, thinking they would go straight through
 - when he actually performed the experiment, he found that the positively charged alpha particles were greatly deflected by the gold (figure 7, p. 104)
 - he determined that the alpha particles must have come very close to another positively charged particle causing them to be repelled and put off course
 - this led him to believe that the positive charge in an atom is not evenly distributed, but concentrated in a very small central area he called the nucleus
 - his model of the atom is a small, positively charged nucleus surrounded by negatively charged particles

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E Durkheim



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Chemistry and Toxicology of Pollution examines the interactions and adverse effects of pollution on both natural ecosystems and human health addressing chemical toxicological and ecological factors at both the regional and global scale The book is written using a conceptual framework that follows the interaction of a pollutant with the environment from distribution in the various abiotic sectors of the environment to exposure and effects on individuals and ecosystems The authors also highlight the critical role of various socio economic political and cultural aspects in achieving sustainable goals strategies and science based solutions to pollution and health This comprehensive volume covers the chemical behavior and governing principles of pollutants their interactions with humans and ecosystems and the methods and processes of environmental risk assessment and pollution management Extensively revised and expanded the second edition equips readers with the knowledge required to help lead the way towards a healthy and sustainable future New chapters address current pollution issues such as global warming and climate change recent advances in environmental science the monitoring and evaluation of new and emerging pollutants risk assessment and remediation and innovative pollution management approaches and techniques With in depth material on human toxicology integrated throughout the text **Chemistry and Toxicology of Pollution** Provides an effective framework for interpreting the information produced by international national and local agencies Presents unifying theories and principles supported by up to date scientific literature Offers broad coverage of pollution science with an emphasis on North America the UK Europe China India and Australia Discusses the similarities and differences of the impact of pollutants on the natural environment and humans **Chemistry and Toxicology of Pollution** Second Edition enables readers to view pollution in its correct perspective and develop appropriate control measures It is essential reading for scientists academic researchers policymakers professionals working in industry and advanced students in need of a clear understanding of the nature and effects of environmental pollution **Prentice Hall Exploring Life Science** Anthea Maton, 1997 **The Chemical News and Journal of Physical Science**, 1907 *Teaching Science for Understanding* James Joseph Gallagher, 2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities Chemical News and Journal of Physical Science, 1887 **Curriculum Review**, 1986 Introduction to Health Physics: Fourth Edition Herman Cember, Thomas E. Johnson, 2008-05-04 A dynamic all inclusive overview of the field of health physics If it s an important topic in the field of health physics you ll find it in this trusted text in sections on physical principles atomic and nuclear structure radioactivity biological effects of radiation and instrumentation This one of a kind guide spans the entire

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