

## Assessment

## Chapter Test B

## Chapter: Chemical Formulas and Chemical Compounds

**PART I** In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. Changing a subscript in a correctly written chemical formula will
- change the electron configuration of that element.
  - change the charges on the other ions in the compound.
  - change the formula so that it no longer represents the same compound.
  - have no effect on the formula.
- \_\_\_\_\_ 2. Using the Stock system of nomenclature,  $\text{Cr}_2(\text{SO}_4)_3$  is named
- chromium(II) sulfate.
  - chromic sulfate.
  - dichromium trisulfate.
  - chromium(III) sulfate.
- \_\_\_\_\_ 3. In a polyatomic ion, the algebraic sum of the oxidation numbers of all atoms is equal to
- 0.
  - 10.
  - the number of atoms in the ion.
  - the charge on the ion.
- \_\_\_\_\_ 4. The first part of the name of a binary ionic compound is the
- cation.
  - polyatomic ion.
  - oxyanion.
  - anion.
- \_\_\_\_\_ 5. The empirical formula may not represent the actual composition of a(n)
- ionic compound.
  - crystal.
  - salt.
  - molecular compound.
- \_\_\_\_\_ 6. The formula for phosphoric acid is
- $\text{HPO}_4$
  - $\text{H}(\text{PO}_4)_2$
  - $\text{H}_2\text{PO}_3$
  - $\text{H}_3\text{PO}_4$

# Chapter Test B Chemical Formulas And Compounds

**American Medical Association**



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