

Chapter - 1

Chemical Reactions and Equations

Chemical Reaction: • Chemical reaction is a process in which new substances with different properties are formed.

- Atoms of one element don't change into another element.
- Only rearrangement of atoms takes place in a chemical reaction.

Examples 1. Hydrogen + Oxygen \rightarrow water

2. Iron + Oxygen \rightarrow rust

3. Potassium and chlorine gas \rightarrow chloride

Physical Change: • No new substance is formed.

- It is a temporary change.
- No change in the composition of the substance takes place.
- It is easily reversible.
- Heat or light may or may not be given out or absorbed.

Chemical Change: • A new substance is formed.

- It is a permanent change.
- The composition of the new substance change greatly.
- The change may be reversible or irreversible.
- Heat or light or both are generally given out or absorbed.

Law of Conservation of mass

Matter can neither be created nor destroyed in a chemical reaction. This means that total mass of all reactants must

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Forensic Chemistry Michael Grossman, 2021-12-20 FORENSIC CHEMISTRY FUNDAMENTALS strives to help scientists workplace health safety and other areas are included And two issues of Canadian legal process are argued as essays in the final two chapters

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