

1. A 33.7-g sample of HF is dissolved in water to give 2.0×10^2 mL of solution. The concentration of the solution is:

- A) 1.68 M
B) 0.34 M
C) 0.17 M
D) 8.4 M
E) 16.9 M

$$[HF] = \frac{33.7g HF}{2.0 \times 10^2 mL} \times \frac{1000 mL}{1 L} \times \frac{1 mol HF}{20.01g HF} = 8.4 M$$

2. What mass of calcium chloride, $CaCl_2$, is needed to prepare 3.419 L of a 1.56 M solution?

- A) 243 g
B) 5.33 g
C) 50.6 g
D) 111 g
E) 592 g

$$\frac{3.419 L}{1} \times \frac{1.56 mol CaCl_2}{L} \times \frac{110.98 g CaCl_2}{1 mol CaCl_2} = 592 g CaCl_2$$

3. A 51.24-g sample of $Ba(OH)_2$ is dissolved in enough water to make 1.205 liters of solution. How many mL of this solution must be diluted with water in order to make 1.000 liter of 0.1000 molar $Ba(OH)_2$?

- A) 403 mL**
B) 248 mL
C) 24.8 mL
D) 83.0 mL
E) 331 mL

$$[Ba(OH)_2] = \frac{51.24g Ba(OH)_2}{1.205 L} \times \frac{1 mol Ba(OH)_2}{171.36g Ba(OH)_2} = 0.2481 M$$

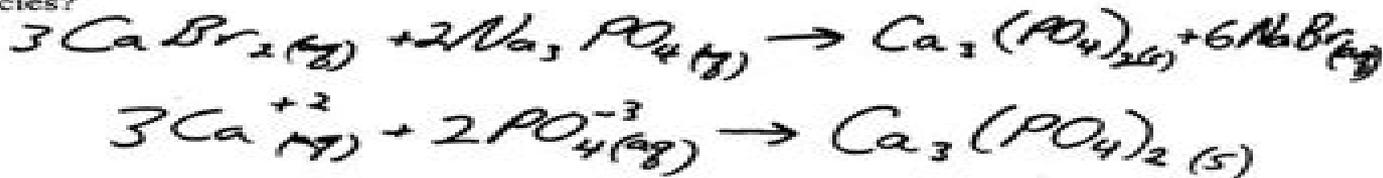
$$M_1 = 0.1000 M \quad M_2 = 0.2481 M$$

$$V_1 = 1.000 L \quad V_2 = ?$$

$$V_2 = \frac{0.1000 M \times 1.000 L}{0.2481 M} = 0.4031 L = 403 mL$$

4. The net ionic equation for the reaction of calcium bromide and sodium phosphate contains which of the following species?

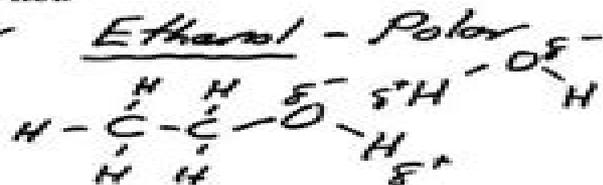
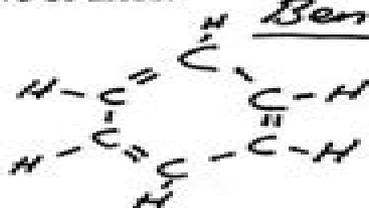
- A) $Ca^{2+}(aq)$
B) $PO_4^{3-}(aq)$
C) $2Ca_3(PO_4)_2(s)$
D) $6NaBr(aq)$
E) $3Ca^{2+}(aq)$



5. Consider two organic molecules, ethanol and benzene. One dissolves in water and the other does not. Why?

- A) They have different molar masses.
B) One is ionic, the other is not.
C) One is an electrolyte, the other is not.
D) Ethanol contains a polar O—H bond, and benzene does not.
E) Two of these.

* Know the general solubility rule of "Likes Dissolve Likes"



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The Technique of Study in First Year Chemistry Edward Wallace Johnson, 1930

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Sandra Smith's Review for NCLEX-PN Sandra Fucci Smith, 2005 Recently Acquired Questions based on Critical Thinking Principles NCLEX Procedures Study Guidelines Test taking Tips and a CD ROM with more than 2300 Q As are included in this all in one resource Topics include Management Principles Legal Issues Nursing Through the Life Cycle General Nursing Concepts Pharmacology Nutritional Management Laboratory Tests Infection Control Disaster Nursing Bioterrorism Medical Surgical Nursing Emergency Nursing Oncology Nursing Geriatric Nursing Maternal Newborn Gynecological Nursing Pediatric Nursing Mental Health Nursing

Holt Chemistry Ralph Thomas Myers, 2004

Study Guide Joan I. Senyk, Darrell D. Ebbing, 1984

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