

Physics Notes Class 11 CHAPTER 12

THERMODYNAMICS

The branch dealing with measurement of temperature is called thermometry and the devices used to measure temperature are called thermometers.

Heat

Heat is a form of energy called thermal energy which flows from a higher temperature body to a lower temperature body when they are placed in contact.

Heat or thermal energy of a body is the sum of kinetic energies of all its constituent particles, on account of translational, vibrational and rotational motion.

The SI unit of heat energy is joule (J).

The practical unit of heat energy is calorie.

$$1 \text{ cal} = 4.18 \text{ J}$$

1 calorie is the quantity of heat required to raise the temperature of 1 g of water by 1°C.

Mechanical energy or work (W) can be converted into heat (Q) by $1 \text{ W} = JQ$

where J = Joule's mechanical equivalent of heat.

J is a conversion factor (not a physical quantity) and its value is 4.186 J/cal.

Temperature

Temperature of a body is the degree of hotness or coldness of the body. A device which is used to measure the temperature, is called a thermometer.

Highest possible temperature achieved in laboratory is about 10⁸ while lowest possible temperature attained is 10⁻⁸ K.

Branch of Physics dealing with production and measurement temperature close to 0 K is known as cryogenics, while that dealing with the measurement of very high temperature is called pyrometry. Temperature of the core of the sun is 10⁷ K while that of its surface 6000 K.

NTP or STP implies 273.15 K (0°C = 32°F).

Different Scale of Temperature

Physics Notes Class 11 Chapter 12 Thermodynamics

Xiaolong Qi



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Table of Contents Physics Notes Class 11 Chapter 12 Thermodynamics

1. Understanding the eBook Physics Notes Class 11 Chapter 12 Thermodynamics
 - The Rise of Digital Reading Physics Notes Class 11 Chapter 12 Thermodynamics
 - Advantages of eBooks Over Traditional Books
2. Identifying Physics Notes Class 11 Chapter 12 Thermodynamics
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physics Notes Class 11 Chapter 12 Thermodynamics
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physics Notes Class 11 Chapter 12 Thermodynamics
 - Personalized Recommendations
 - Physics Notes Class 11 Chapter 12 Thermodynamics User Reviews and Ratings

- Physics Notes Class 11 Chapter 12 Thermodynamics and Bestseller Lists
- 5. Accessing Physics Notes Class 11 Chapter 12 Thermodynamics Free and Paid eBooks
 - Physics Notes Class 11 Chapter 12 Thermodynamics Public Domain eBooks
 - Physics Notes Class 11 Chapter 12 Thermodynamics eBook Subscription Services
 - Physics Notes Class 11 Chapter 12 Thermodynamics Budget-Friendly Options
- 6. Navigating Physics Notes Class 11 Chapter 12 Thermodynamics eBook Formats
 - ePub, PDF, MOBI, and More
 - Physics Notes Class 11 Chapter 12 Thermodynamics Compatibility with Devices
 - Physics Notes Class 11 Chapter 12 Thermodynamics Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physics Notes Class 11 Chapter 12 Thermodynamics
 - Highlighting and Note-Taking Physics Notes Class 11 Chapter 12 Thermodynamics
 - Interactive Elements Physics Notes Class 11 Chapter 12 Thermodynamics
- 8. Staying Engaged with Physics Notes Class 11 Chapter 12 Thermodynamics
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physics Notes Class 11 Chapter 12 Thermodynamics
- 9. Balancing eBooks and Physical Books Physics Notes Class 11 Chapter 12 Thermodynamics
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physics Notes Class 11 Chapter 12 Thermodynamics
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physics Notes Class 11 Chapter 12 Thermodynamics
 - Setting Reading Goals Physics Notes Class 11 Chapter 12 Thermodynamics
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physics Notes Class 11 Chapter 12 Thermodynamics
 - Fact-Checking eBook Content of Physics Notes Class 11 Chapter 12 Thermodynamics
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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