

Subject	Topic	Test - KAR	Date
C + M + P	Combined Test	JEE Adv - 11 - CT JALL20150928	28 <sup>th</sup> Sep 2015

**C+M+P Key Answers:**

1. a	2. d	3. b	4. d	5. a	6. c	7. 5	8. 3	9. 5	10. 5
11. 5	12. b	13. b	14. c	15. b	16. a	17. d	18.	19.	20. a
21. d	22. c	23. a	24. b	25. c	26. 1	27. 3	28. 3	29. 1	30. 4
31. c	32. a	33. b	34. c	35. b	36. a	37.	38.	39. c	40. a
41. c	42. c	43. a	44. b	45. 4	46. 1	47. 5	48. 1	49. 4	50. b
51. c	52. d	53. b	54. d	55. b	56.	57.			

18. A - p, B - p,q, C - r, D - p,s ;

37. A - q, B - r, C - s, D - s,p ;

56. A - s, B - r, C - p, D - q ;

19. A - q, B - r, C - s, D - p ;

38. A - r, B - s, C - p, D - q ;

57. A - r, B - p, C - q, D - t ;

**Chemistry Solutions:**

- 2-hydroxypropane-1,2,3-tricarboxylic acid
- 2-methylpentane. In 2 position 1 methyl group is there so these compound is having isopropyl group
- 1-Hydroxy - 7 - methyl spiro ( 2,5) oct - 4 - en - 6 - one
- It is white in colour
- 20ml. dilute unreacted acid soln. required

$$= 32 \text{ mL } \frac{N}{10} \text{ NaOH soln.}$$

500 ml. dilute unreacted acid soln. required

$$= \frac{32}{20} \times 500 \text{ mL } \frac{N}{10} \text{ NaOH soln.}$$

$$= \frac{32}{20} \times \frac{500}{10} \text{ mL } N \text{ NaOH}$$

$$= 80 \text{ mL } N \text{ NaOH}$$

$$80 \text{ mL } N \text{ NaOH} = 80 \text{ mL } N \text{ H}_2\text{SO}_4$$

Acid used for the neutralization of  $\text{NH}_3$ 

$$= (100 - 80) \text{ mL } N \text{ H}_2\text{SO}_4$$

$$= 20 \text{ mL } N \text{ H}_2\text{SO}_4$$

$$\text{Percentage of nitrogen} = 1.4 \times N_1 \times \frac{V}{W}$$

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**ML Yell**



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