



3: Chemical Kinetics

Standard: 12th Science

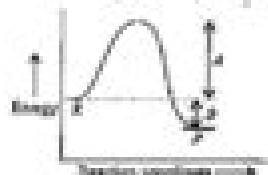
Subject: Chemistry

Student Name: _____

Q1. Consider the graph given in Fig. Which of the following options does not show instantaneous rate of reaction at 40th second?

1. $\frac{V_2 - V_1}{50 - 30}$
2. $\frac{V_1 - V_2}{50 - 30}$
3. $\frac{V_2 - V_1}{30 - 20}$
4. $\frac{V_1 - V_2}{30 - 20}$

Q2. The energy diagram of a reaction $P + Q \rightarrow R + S$ is given. What are A and B in the graph?



1. A \rightarrow activation energy, B \rightarrow heat of reaction
 2. A \rightarrow threshold energy, B \rightarrow heat of reaction
 3. A \rightarrow heat of energy, B \rightarrow activation reaction
 4. A \rightarrow potential energy, B \rightarrow energy of reaction
- Q3.** Decomposition of nitrogen pentoxide is known to be a first order reaction 75 percent of the oxide had decomposed in the first 24 minutes. At the end of an hour, after the start of the reaction, the amount of oxide left will be
1. Nil
 2. About 1%
 3. About 2%
 4. About 3%
- Q4.** Higher order (>3) reactions are rare due to:
1. Low probability of simultaneous collision of all the reacting species.
 2. Increase in entropy and activation energy as more molecules are involved.



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