

AP Chemistry Final Study Guide

Unit 1: Atomic Structure and Properties (7%–9%)

- At STP, one mole of gas always occupies 22.4 L
- **Mass Spectroscopy:**
 - Atoms have different isotopes (carbon-12, carbon-14, etc) and this will affect the average atomic mass
 - The most abundant isotope will have the greatest effect
- **Coulomb's Law:**
 - The greater the charge of the nucleus, the more energy is required to ionize an electron
 - The closer an electron is to the nucleus, the more energy is required to ionize an electron
- **Electromagnetic Radiation:**
 - When atoms absorb energy, electrons jump to higher energy levels
 - When electrons drop from higher to lower energy levels, atoms give off electromagnetic radiation energy
- **Kinetic Energy of Ionized Electrons:**
 - The faster an ejected electron is going, the more kinetic energy it has
 - Electrons originally further away from the nucleus require less energy to eject, so it will have more kinetic energy
- **Electron Configuration:**
 - When entering the d-block (transition metals), the energy level drops by 1 unless you are using shorthand notation
 - Nickel: $1s^2 2s^2 2p^6 3s^2 3p^6 3d^8 4s^2$ (notice the d-block)
 - Nickel (with shorthand): $[\text{Ar}]4s^2 3d^8$
- **Law of Definite Proportions:**
 - A pure chemical compound broken down into elements always contains elements of a fixed ratio, independent of where and how it was created
- **Law of Multiple Proportions:**
 - When elements form compounds, the proportions of the elements in those chemical compounds can be expressed in small whole number ratios
- **Aufbau Principle:**
 - When building up the electron configuration of an atom, electrons are placed in orbitals, subshells, and shells in order of increasing energy

Ap Chemistry Exam Study Guide

DJ Losen



Ap Chemistry Exam Study Guide:

Ap Chemistry Exam Study Guide Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Ap Chemistry Exam Study Guide**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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