

2018

Name: Phyllis A. Smith Lab Section: 2018

Suppose that you had the correct apparatus for the following experiment. In what state would the liquid be at room temperature?

CH<sub>2</sub>Br<sub>2</sub>(l)

CH<sub>2</sub>Br<sub>2</sub>(g)

CH<sub>2</sub>Br<sub>2</sub>(s)

CH<sub>2</sub>Br<sub>2</sub>(aq)

CH<sub>2</sub>Br<sub>2</sub>(l)

Correct Answer

CH<sub>2</sub>Br<sub>2</sub>(l)

CH<sub>2</sub>Br<sub>2</sub>(l)

1. If you had the same conditions as a 1.00 g sample of CH<sub>2</sub>Br<sub>2</sub> at 25 °C, how much could be condensed by compression of water to yield a 1.00 g sample of water?

$$\frac{1.00 \text{ g } \text{CH}_2\text{Br}_2(l)}{100 \text{ g}} \left| \frac{1 \text{ mol}}{253 \text{ g}} \right| \frac{1 \text{ mol}}{1 \text{ mol}} \left| \frac{18 \text{ g}}{1 \text{ mol}} \right| = 0.139 \text{ mol } \text{H}_2\text{O}(l)$$

0.139 mol

2. If you were in the market and wanted the same amount of CH<sub>2</sub>Br<sub>2</sub> at 25 °C, how much of the liquid would be needed at 100 °C?

$$1.00 \text{ g } \text{CH}_2\text{Br}_2(l) \left| \frac{1 \text{ mol}}{100 \text{ g}} \right| \frac{1 \text{ mol}}{1 \text{ mol}} \left| \frac{18 \text{ g}}{1 \text{ mol}} \right| = 0.182 \text{ mol } \text{H}_2\text{O}(l)$$

$$1.00 \text{ g } \text{CH}_2\text{Br}_2(l) \left| \frac{1 \text{ mol}}{100 \text{ g}} \right| \frac{1 \text{ mol}}{1 \text{ mol}} \left| \frac{18 \text{ g}}{1 \text{ mol}} \right| = 0.182 \text{ mol } \text{H}_2\text{O}(l)$$

0.182 mol

Answer

0.182 mol

Answer

# Chem 110 Lab Questions And Answers

**Robert R. Redfield**



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