

Test "Chemistry"

Chapter # 3

Objective part

Time: 15 min

Total marks: 12

Q.1- Every question has four possible answers. Encircle the correct option.

- i) The atomic radii of the elements in periodic table;
a) increase from left to right in a period
b) increase from top to bottom in a group
c) do not change from left to right in a period
d) decrease from top to bottom in a group
- ii) Mendeleev Periodic table was based upon the:
a) electronic configuration
b) atomic mass
c) atomic number
d) electro negativity
- iii) The amount of energy given out when an electron is added to an atom is called;
a) ionization energy
b) lattice energy
c) electron affinity
d) electro negativity
- iv) Long form of periodic table is constructed on the basis of:
a) Mendeleev postulate
b) atomic number
c) atomic mass
d) mass number
- v) Along the period which of the following decreases:
a) electro negativity
b) ionization energy
c) atomic radius
d) electron affinity
- vi) Mark the incorrect statement about the ionization energy:
a) it is measured kJmol^{-1}
b) it is the absorption of energy
c) it decreases in a period
d) it decreases in groups
- vii) Transition elements are:
a) all gases
b) all metals
c) all non metals
d) all metalloids
- viii) S-f series of inner transition elements are called:
a) lanthanides
b) actinides
c) halogens
d) alkali metals
- ix) If we move from left to right in a period the value of ionization energy:
a) remain same
b) decreases
c) increases
d) not effected
- x) Which scientist has given the idea of actives in periodic table:
a) Mendeleev
b) Al-razi
c) Newland
d) dobernier
- xi) Shielding effect is greater in atoms with greater number of:
a) proton
b) electron
c) neutron
d) positron
- xii) Which is the incomplete period in the periodic table:
a) 4th period
b) 2nd period
c) 7th period
d) 1st period

Chemistry Chapter 3 Test Answers

Robert R. Redfield



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Table of Contents Chemistry Chapter 3 Test Answers

1. Understanding the eBook Chemistry Chapter 3 Test Answers
 - The Rise of Digital Reading Chemistry Chapter 3 Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 3 Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 3 Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 3 Test Answers
 - Personalized Recommendations
 - Chemistry Chapter 3 Test Answers User Reviews and Ratings
 - Chemistry Chapter 3 Test Answers and Bestseller Lists
5. Accessing Chemistry Chapter 3 Test Answers Free and Paid eBooks

- Chemistry Chapter 3 Test Answers Public Domain eBooks
 - Chemistry Chapter 3 Test Answers eBook Subscription Services
 - Chemistry Chapter 3 Test Answers Budget-Friendly Options
6. Navigating Chemistry Chapter 3 Test Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Chemistry Chapter 3 Test Answers Compatibility with Devices
 - Chemistry Chapter 3 Test Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chemistry Chapter 3 Test Answers
 - Highlighting and Note-Taking Chemistry Chapter 3 Test Answers
 - Interactive Elements Chemistry Chapter 3 Test Answers
8. Staying Engaged with Chemistry Chapter 3 Test Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 3 Test Answers
9. Balancing eBooks and Physical Books Chemistry Chapter 3 Test Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 3 Test Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 3 Test Answers
- Setting Reading Goals Chemistry Chapter 3 Test Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 3 Test Answers
- Fact-Checking eBook Content of Chemistry Chapter 3 Test Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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