

$$\text{a) Rate} = \frac{-1}{2} \frac{d[\text{NO}]}{dt} = \frac{-d[\text{O}_2]}{dt} = \frac{1}{2} \frac{d[\text{NO}_2]}{dt}$$

$$\text{b) } \frac{-d[\text{O}_2]}{dt} = \frac{1}{2} \frac{d[\text{NO}_2]}{dt}$$

$$\frac{d[\text{NO}_2]}{dt} = 2 \times \left(\frac{-d[\text{O}_2]}{dt} \right) = 2 \times 0.2 \text{ mol L}^{-1} \text{ s}^{-1}$$

$$= 0.4 \text{ mol L}^{-1} \text{ s}^{-1}$$

Chemistry Chemical Kinetics Problems Answers

Lingsheng Yao



Chemistry Chemical Kinetics Problems Answers:

The Enigmatic Realm of **Chemistry Chemical Kinetics Problems Answers**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing in short supply of extraordinary. Within the captivating pages of **Chemistry Chemical Kinetics Problems Answers** a literary masterpiece penned with a renowned author, readers attempt a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting effect on the hearts and minds of those who partake in its reading experience.

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