

Lab Data

Verify your calculation. In your set-up, do your units cancel?
Did you report your data to the correct number of significant figures?


Moles of magnesium (mol)	0.00049
Temperature of water (°C)	22.0
Temperature of water (K)	295.15
Vapor pressure of water (mmHg)	19.8
Barometric pressure (mmHg)	776.0

Observations

bubbles evident inside the eudiometer

Volume of hydrogen gas collected (mL)	12.00
Volume of hydrogen gas collected (L)	0.01200
Pressure of hydrogen gas (mmHg)	760.4
Pressure of hydrogen gas (atm)	1.0005265
Moles of hydrogen gas produced (mol)	0.00049716
Ideal gas constant (L·atm·mol ⁻¹ ·K ⁻¹)	0.0821

 [How to calculate ideal gas constant](#)

 [Water Vapor Pressure Table](#)

Ideal Gas Constant Lab 38 Answers

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**

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Chemistry Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02

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