

Buffer Solution

The diagram illustrates the effect of adding acid (HCl) to two different solutions. In the top row, a pH meter with a reading of 7.0 is shown in a beaker of water (H_2O) on the left and a beaker of water with a buffer ($H_2O + \text{Buffer}$) on the right. A box labeled "Adding acid (HCl)" has arrows pointing down to the bottom row. In the bottom row, the left beaker now contains a pink acidic solution with a pH meter reading of 3.0, labeled "pH lowers, turns acidic". The right beaker still contains a blue solution with a pH meter reading of 6.9, labeled "Stable pH".

pH meter
Electrode

H_2O $H_2O + \text{Buffer}$

Adding acid (HCl)

3.0 **6.9**

pH lowers, turns acidic **Stable pH**

The diagram illustrates the effect of adding base (NaOH) to two different solutions. In the top row, a pH meter with a reading of 7.0 is shown in a beaker of water (H_2O) on the left and a beaker of water with a buffer ($H_2O + \text{Buffer}$) on the right. A box labeled "Adding base (NaOH)" has arrows pointing down to the bottom row. In the bottom row, the left beaker now contains a dark blue basic solution with a pH meter reading of 11.0, labeled "pH rises, turns basic". The right beaker still contains a blue solution with a pH meter reading of 7.1, labeled "Stable pH".

pH meter
Electrode

H_2O $H_2O + \text{Buffer}$

Adding base (NaOH)

11.0 **7.1**

pH rises, turns basic **Stable pH**

Preparation And Properties Of Buffer Solutions Pre Lab Answers

Michael Seilmaier



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