

# Complete Guide to Atomic Structure and the Periodic Table

## Chapter 1: Introduction to Atoms

### What is an Atom?

An atom is the smallest particle into which an element can be divided while still maintaining the properties of that element. Think of it as the fundamental building block of matter. To understand this concept, imagine cutting a gold bar in half repeatedly. Eventually, you would reach a point where cutting it further would no longer leave you with gold - that smallest piece that still retains the properties of gold is an atom.

### The Three Components of an Atom

Every atom consists of three fundamental particles:

#### 1. Protons (Positive Charge)

- Located in the nucleus (center) of the atom
- Carry a positive electrical charge (+)
- Determine the identity of the element

#### 2. Neutrons (Neutral Charge)

- Also located in the nucleus
- Have no electrical charge (neutral)
- Add mass to the atom without affecting its chemical identity

#### 3. Electrons (Negative Charge)

- Orbit around the nucleus in energy levels or shells
- Carry a negative electrical charge (-)
- Much smaller in mass compared to protons and neutrons

### The Bohr Model of the Atom

The Bohr model, while simplified, provides an excellent foundation for understanding atomic structure. In this model:

- The nucleus contains protons and neutrons at the center
- Electrons orbit the nucleus in defined energy levels (shells)
- The first shell can hold up to 2 electrons
- The second shell can hold up to 8 electrons
- The third shell can hold up to 8 electrons (in simple cases)

# Atomic Structure And Periodic Relationships Study Guide

**J Dewey**



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