

PHYSICAL PHARMACY

PHYSICAL CHEMISTRY – branch of chemistry concerned with physical properties and structure of matter and with laws and theories of physical and chemical changes

PHYSICAL PHARMACY – application of physical chemistry in pharmacy; study of physicochemical properties of substances used in drug formulation

I. FORCES OF ATTRACTION

A. INTRAMOLECULAR – intra (within) the molecule

- **Ionic Bond** – transfer of electrons; metal (electron donor) + non-metal (electron acceptor); (ex. Na: $p^6 = 11$; $e^- = 11 \rightarrow p^7 = 11$; $e^- = 10 \rightarrow$ charge = +1)

○ PERIODIC TRENDS REPRESENTING BONDS

- **Ionization Energy** – energy required remove electron (low M; high NM)
- **Electronegativity** – ability to attract electrons (low M; high NM)

- **Covalent Bonds** – sharing of electrons; non-metal + non-metal; polarity – difference in electronegativity (symmetrical = non-polar)

- **Polar** – unequal sharing of electrons; dipole moment – (ex. $H^{\delta+}-Br^{\delta-}$)
- **Non Polar** – equal sharing of electrons (ex. $Br-Br$)

B. INTERMOLECULAR – inter (between) the molecules

- **Van der Waals**

- **Dipole-Dipole** → P + P; Keesom Forces; orientation effect
- **Dipole-Induced Dipole** → P + NP; Debye Forces; inductive effect
- **Induced Dipole-Induced Dipole** → NP + NP; London Forces; dispersive effect

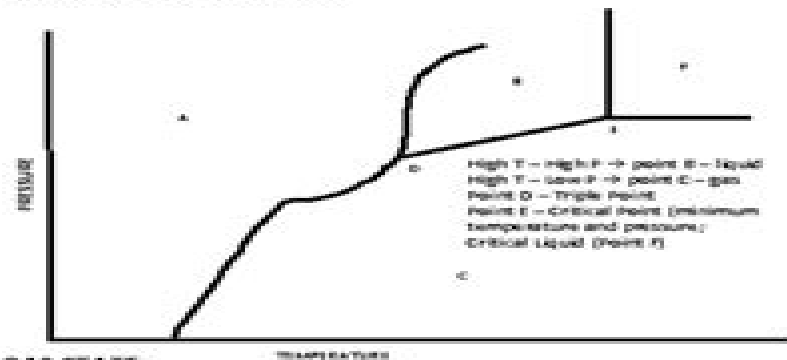
- **Ion-Induced Dipole** → ion + NP

- $I_2 \rightarrow NP$ (solubility = 1:2960)
- $I_2 + NaI/KI$ (solubilizing agent) → $I_3 + I_2$ = increase solubility

- **Hydrogen Bond** – hydrogen atom + electronegative atom (N, O, F)



PHASE DIAGRAM



GAS STATE:

- **Kinetic Molecular Theory**
 - Gases are composed of small particles called atoms; total volume is negligible in relation to space where they are confined
 - Particles do not attract one another but move with complete independence
 - Particles exhibit continuous random motion owing to their kinetic energy
 - Gases exhibit perfect elasticity (no energy exchange)

II. GAS LAWS

- at standard temperature (Kelvin = $^{\circ}C + 273$) and pressure (1 atm = 760 mmHg or torr; 760 mmHg; 1.01325×10^5 dynes/cm²; 1.01325×10^5 Pa or N/m²; 0.1325 bar)
- **Boyle's Law** – constant temperature; increase P = decrease V; inversely related

$$P \propto \frac{1}{V}$$

- * density = increase altitude; decrease volume; $P_1V_1 = P_2V_2$

- **Charles' Law** – constant pressure; increase T = increase V

$$T \propto V; \frac{T_1}{T_2} = \frac{V_1}{V_2} \text{ or } \frac{T_1}{V_1} = \frac{T_2}{V_2}$$

- * H_2O – 1.0g/ml → $4^{\circ}C$

- **Gay Lussac's Law** → constant volume

$$P \propto T; \frac{P_1}{P_2} = \frac{T_1}{T_2} \text{ or } \frac{P_1}{T_1} = \frac{P_2}{T_2}$$

- * $121^{\circ}C$ – 15-20psi

Physical Pharmacy Lecture Notes

**Temple University. School of
Pharmacy**



Physical Pharmacy Lecture Notes:

Physical Pharmacy Lecture Notes Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Physical Pharmacy Lecture Notes**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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