

Chapter 25: NUCLEAR CHEMISTRY

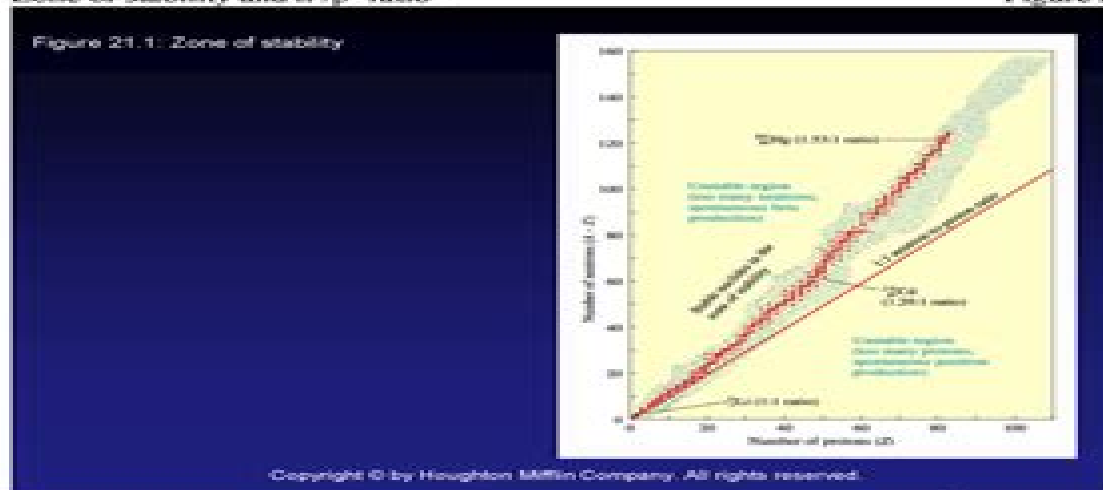
The number and arrangement of its electrons dominate the chemical reactivity and physical properties of an atom.

Hence, we focus on electron configurations in chemistry.

However, we live in a world with nuclear power plants, nuclear weapons, and medical applications of radioactive isotopes, including waste from their use and production.

Nuclear radiation can break chemical bonds, including bonds in DNA.

§25-7 Nuclear stability

Zone of stability and n^0/p^+ ratioFigure 25-7: Zumdahl 4th Fig 21.1

neutrons and protons are attracted to each other, so the more neutrons, [What?]

p^+ = proton

n^0 = neutron

e^- = electron

Stable nuclei most likely have even number of p^+ & even number of n^0 (Table 25-3)

Number of p^+	Number of n^0	Number of naturally occurring stable nuclides	Examples
even	even	163	^4He , ^{12}C , ^{16}O , ^{18}O , ^{20}Ne
even	odd	55	^3He , ^7Be , ^{13}C , ^{17}O , ^{21}Ne
odd	even	50	^7Li , ^{11}B , ^{14}N , ^{19}F , ^{23}Na
odd	odd	4	^2H , ^6Li , ^{10}B , ^{14}N

Shell model for nucleus gives magic numbers: 2, 8, 20, 28, 50, 82, 114 (for p^+), 126 (for n^0)

<http://www.nndc.bnl.gov/nudat2/>

Brookhaven National Laboratory puts # neutrons on x axis

§25-1 Nuclear reactions destroy atoms -- contrary to Dalton's Atomic Theory
But Mass number = # nucleons (p^+ & n^0) is unchanged

\Rightarrow mass is approximately conserved

Mass-energy is conserved. $\Delta E = (\Delta m)c^2$

Charge is conserved.

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S Marginson



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