

Chapter 6 Thermochemistry
Homework Problems

- The enthalpy change for the fermentation (without oxygen, O_2 gas) reaction of glucose to alcohol and carbon dioxide is -67 kJ.

 - Write a chemical equation for the fermentation reaction and show how energy is involved in the reaction.
 - Is the reaction exothermic or endothermic?
 - Is energy, in the form of heat, absorbed or evolved as the reaction occurs?
- The specific heat capacity of graphite is 0.71 J/°C·g. Calculate the energy required to:

 - Raise the temperature of 1.0 mol graphite by 1.0 °C.
 - Raise the temperature of 850 g graphite by 150 °C.
 - Raise the temperature of 75 kg graphite from 294 K to 348 K.
- In a coffee-cup calorimeter, 100.0 mL of 1.0 M NaOH and 100.0 mL of 1.0 M HCl are mixed. Both solutions were originally at 24.6 °C. After the reaction, the final temperature is 31.3 °C. Assume that all the solutions have a density of 1.0 g/mL and a specific heat capacity of 4.18 J/g°C. Calculate the enthalpy change for the neutralization of HCl by NaOH. Assume that no heat is lost to the surroundings or to the calorimeter.
- Camphor ($C_{15}H_{26}O$) has an energy of combustion of -5903.6 kJ/mol. A sample of camphor with a mass of 0.1204 g is burned in a bomb calorimeter, the temperature increases by 2.28 °C. Calculate the heat capacity of the calorimeter.

A 0.1964 g sample of quinone ($C_6H_4O_2$) is burned in a bomb calorimeter that has a heat capacity of 1.56 kJ/°C. The temperature of the calorimeter increases by 3.2 °C. Calculate the energy of combustion of quinone per gram and per mole.

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Table of Contents Chapter 6 Thermochemistry Homework Problems

1. Understanding the eBook Chapter 6 Thermochemistry Homework Problems
 - The Rise of Digital Reading Chapter 6 Thermochemistry Homework Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 6 Thermochemistry Homework Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 6 Thermochemistry Homework Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 6 Thermochemistry Homework Problems
 - Personalized Recommendations
 - Chapter 6 Thermochemistry Homework Problems User Reviews and Ratings
 - Chapter 6 Thermochemistry Homework Problems and Bestseller Lists
5. Accessing Chapter 6 Thermochemistry Homework Problems Free and Paid eBooks
 - Chapter 6 Thermochemistry Homework Problems Public Domain eBooks
 - Chapter 6 Thermochemistry Homework Problems eBook Subscription Services
 - Chapter 6 Thermochemistry Homework Problems Budget-Friendly Options

6. Navigating Chapter 6 Thermochemistry Homework Problems eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 6 Thermochemistry Homework Problems Compatibility with Devices
 - Chapter 6 Thermochemistry Homework Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 6 Thermochemistry Homework Problems
 - Highlighting and Note-Taking Chapter 6 Thermochemistry Homework Problems
 - Interactive Elements Chapter 6 Thermochemistry Homework Problems
8. Staying Engaged with Chapter 6 Thermochemistry Homework Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 6 Thermochemistry Homework Problems
9. Balancing eBooks and Physical Books Chapter 6 Thermochemistry Homework Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 6 Thermochemistry Homework Problems
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 6 Thermochemistry Homework Problems
 - Setting Reading Goals Chapter 6 Thermochemistry Homework Problems
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 6 Thermochemistry Homework Problems
 - Fact-Checking eBook Content of Chapter 6 Thermochemistry Homework Problems
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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