

## Chapter 8 Solutions, Acids, and Bases

**Section 8.2 Solubility and Concentration**

(pages 235–239)

*This section explains solubility, the factors affecting solubility, and different ways of expressing the concentration of a solution.***Reading Strategy (page 235)**

**Previewing** Before you read the section, rewrite the topic headings as *how*, *why*, and *what* questions. As you read, write an answer to each question. For more information on this reading strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

Question	Answer
What is solubility?	
	Solvent, temperature, and pressure

**Solubility (pages 235–237)**

- Define solubility. \_\_\_\_\_
- List the following solutes in order from most soluble to least soluble in water: table salt, baking soda, table sugar.
  - \_\_\_\_\_
  - \_\_\_\_\_
  - \_\_\_\_\_
- Circle the letters that identify how solutions can be classified based on solubility.
 

a. unsaturated	b. desaturated
c. saturated	d. supersaturated
- Describe a saturated solution. \_\_\_\_\_
- A solution that has less than the maximum amount of solute that can be dissolved is called a(n) \_\_\_\_\_.
- Is the following sentence true or false? It is impossible for a solution to contain more solute than the solvent can hold at a given temperature. \_\_\_\_\_

# Chapter 8 Solutions Section 3 Solubility And Concentration

**P.G.T. Fogg**



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