

## Energy and Chemical Change

### Section 15.1 Energy

pages 516–522

#### Practice Problems

pages 519–521

1. A fruit and oatmeal bar contains 142 nutritional Calories. Convert this energy to calories.

$$142 \text{ Calories} = 142 \text{ kcal}$$

$$142 \text{ kcal} \times \frac{1000 \text{ cal}}{1 \text{ kcal}} = 142,000 \text{ cal}$$

2. An exothermic reaction releases 86.5 kJ. How many kilocalories of energy are released?

$$86.5 \text{ kJ} \times \frac{1 \text{ kcal}}{4.184 \text{ kJ}} = 20.7 \text{ kcal}$$

3. **Challenge** Define a new energy unit, named after yourself, with a magnitude of one-tenth of a calorie. What conversion factors relate this new unit to joules? To Calories?

$$\text{Unit X} = 0.1 \text{ cal}$$

$$1 \text{ cal} = 4.184 \text{ J}$$

$$\text{X} = (0.1 \text{ cal})(4.184 \text{ J/cal}) = 0.4184 \text{ J}$$

$$1 \text{ cal} = 0.001 \text{ Calorie}$$

$$\text{X} = (0.1 \text{ cal})(1 \text{ Cal}/1000 \text{ cal}) = 0.0001 \text{ Calorie}$$

4. If the temperature of 34.4 g of ethanol increases from 25.0°C to 78.8°C, how much heat has been absorbed by the ethanol? Refer to Table 15.2.

$$q = c \times m \times \Delta T$$

$$q = 2.44 \text{ J/(g}\cdot\text{°C)} \times 34.4 \text{ g} \times 53.8\text{°C} = 4.52 \times 10^3 \text{ J}$$

5. A 155-g sample of an unknown substance was heated from 25.0°C to 40.0°C. In the process, the substance absorbed 5696 J of energy. What is the specific heat of the substance? Identify the substance among those listed in Table 15.2 on page 520.

$$q = c \times m \times \Delta T$$

$$c = \frac{q}{m\Delta T} = \frac{(5696 \text{ J})}{(155 \text{ g})(40.0 - 25.0\text{°C})} = 2.45 \text{ J/(g}\cdot\text{°C)}$$

The specific heat is very close to the value for ethanol.

6. **Challenge** A 4.50-g nugget of pure gold absorbed 276 J of heat. The initial temperature was 25.0°C. What was the final temperature?

$$q = c \times m \times \Delta T$$

$$\Delta T = \frac{q}{cm} = \frac{(276 \text{ J})}{(0.129 \text{ J/g}\cdot\text{°C})(4.50 \text{ g})} = 475\text{°C}$$

$$\Delta T = T_f - T_i = T_f - 25.0\text{°C} = 475\text{°C}$$

$$T_f = 5.00 \times 10^2\text{°C}$$

### Section 15.1 Assessment

page 522

7. **Explain** how energy changes from one form to another in an exothermic reaction. In an endothermic reaction.

Chemical potential energy changes to heat in exothermic reactions and the heat is released. In endothermic reactions, heat is absorbed and changed to chemical potential energy.

8. **Distinguish** between kinetic and potential energy in the following examples: two separated magnets; an avalanche of snow; books on library shelves; a mountain stream; a stock-car race; separation of charge in a battery.

Two separated magnets illustrate potential energy. In a snow avalanche, positional potential energy is changing to kinetic energy. Books on a shelf illustrate positional potential energy. As water races down a mountain stream, positional potential energy is changing to kinetic energy. In a stock-car race, chemical potential energy is being changed to kinetic energy. The separation of charge in a battery illustrates electrical potential energy.

9. **Explain** how the light and heat of a burning candle are related to chemical potential energy.

Chemical potential energy, contained in the candle, is changed to energy in the form of light and heat and released as the chemical combustion reaction takes place.

# Chemistry Matter Change Chapter 15 Answer Key

**McGraw Hill**



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**Solutions Manual for Chemistry: Molecules Matter and Change, Fourth Edition** Julie Henderleiter, Charles Trapp, Peter Atkins, Loretta Jones, 1999-08-13 This student companion is a supplement to Chemistry Molecules Matter and Change 4th edition with CD ROM It features guided reading strategies collaborative learning sheets and strategies for using CD ROM tools Chemistry: Matter & Change, Solving Problems - A Chemistry Handbook McGraw Hill, 2001-08 Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change **Ebook: Chemistry: The Molecular Nature of Matter and Change** Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change **Chemistry** McGraw-Hill Staff, 2001-03 **Chemistry: Science of Matter, Energy, and Change** Gregory R. Choppin, Bernard Jaffe, 1965 *Chemistry* James E. Brady, John R. Holm, 1993-01-21 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples Chemistry E. Russell Hardwick, 1965 **Cehmistry Textbook for College and University USA** Ibrahim Sikder, 2023-06-04 Cehmistry Textbook USA *Norman Hall's Asvab Preparation Book* Norman Hall, 2015-01-02 Everything you need to know thoroughly covered in one book five ASVAB practice tests answer keys tips to boost scores military enlistment information study aids **A Text-book of Physics and Chemistry for Nurses** Andrew Richard Bliss, Alfred Henderson Olive, 1923 Chemical News and Journal of Industrial Science, 1902 *The Chemical News and Journal of Physical Science*, 1911 Journal of the Society of Chemical Industry Society of Chemical Industry (Great Britain), 1888 Includes list of members 1882 1902 and proceedings of the annual meetings and various supplements Structure and Change Gordon Sechrist Christiansen, Paul Hobart Garrett, 1960 *Chiropractic, Other Drugless Healing Methods* Robert Emmette McNamara, 1913 **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 A Course in inorganic chemistry for colleges Lyman Churchill Newell, 1909 Climate Change and Estuaries Michael J. Kennish, Hans W. Paerl, Joseph R. Crosswell, 2023-09-15 Climate change is having an increasing impact on coastal estuarine and marine environments worldwide This book provides state of the art coverage of climate change effects on estuarine ecosystems from local regional and global perspectives With editors among the most noted international scholars in coastal ecology and estuarine science and contributors who are world class in their fields the chapters in this volume consist of comprehensive studies in coastal estuarine and marine sciences climate change and coastal management and provide an extensive international collection of data in tabular illustrated and narrative formats useful for coastal scientists planners and managers Comprised of three sections 1 physical chemical aspects 2 biological aspects and 3 management aspects the book not only examines climatic and non climatic drivers of

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