

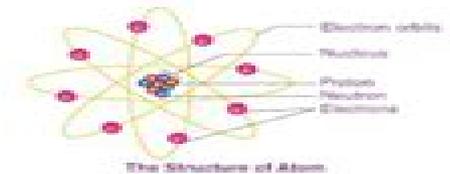
Chapter 2: The Chemistry of Life

➤ Legos of the Universe

- Atoms
- Elements
 - Gold, carbon, silver, etc. (periodic table)

➤ Atomic Structure

- Nucleus
 - Proton (+)
 - Neutrons (no charge)
- Electron Shell



➤ Isotopes

- Atomic Mass

➤ Radioactive Atoms

- Some atomic nuclei are not stable and break down randomly
- These atoms are radioactive
- During decomposition they release, at a constant rate, a tiny high-speed particle carrying a lot of energy

➤ Electron shells determine bonding patterns

➤ Versatility of Carbon

- Can make 4 bonds
- Most often with C, H, O, and N
- Huge potential variety

➤ Ions

- Particles in the outer shell

➤ Changing Parts of the Atom

- What happens if you...
 - Add an electron to chlorine
 - Add a neutron to nitrogen
 - Add a proton to hydrogen



Number of protons = 11

Number of electrons = 11

Charge = 0 (Neutral)

Number of protons = 11

Number of electrons = 10

Charge = +1

➤ Elements in Living Things

- Oxygen
- Carbon
- Hydrogen
- Nitrogen

➤ Molecules

- Molecules
- Bond energy
- Chemical reactions
 - **Reactants > Products**

➤ Chemical Bonding

- Covalent
- Ionic

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N Noddings



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