

**Chemistry: Measurement: Problem Sheet 2**

**Warning:** There are 100 marks available for this problem sheet. Please plan your time accordingly (approximately 100 minutes).

1. Mass spectrometry experiments according to the following reactions sequence:



(A) How many moles of gas are needed to produce 100 moles of  $\text{C}_2\text{H}_5\text{O}^+$ ?

$$\text{C}_2\text{H}_6\text{O} + \text{H}^+ \rightarrow \text{C}_2\text{H}_7\text{O}^+ \rightarrow \text{C}_2\text{H}_5\text{O}^+ + \text{H}_2 \rightarrow \text{C}_2\text{H}_5^+ + \text{H}_2\text{O}$$

(B) How many moles of gas are needed to produce 100 moles of  $\text{C}_2\text{H}_5^+$ ?

$$\text{C}_2\text{H}_6\text{O} + \text{H}^+ \rightarrow \text{C}_2\text{H}_7\text{O}^+ \rightarrow \text{C}_2\text{H}_5\text{O}^+ + \text{H}_2 \rightarrow \text{C}_2\text{H}_5^+ + \text{H}_2\text{O}$$

(C) How many moles of gas are needed to produce 100 moles of  $\text{C}_2\text{H}_7\text{O}^+$ ?

$$\text{C}_2\text{H}_6\text{O} + \text{H}^+ \rightarrow \text{C}_2\text{H}_7\text{O}^+ \rightarrow \text{C}_2\text{H}_5\text{O}^+ + \text{H}_2 \rightarrow \text{C}_2\text{H}_5^+ + \text{H}_2\text{O}$$

(D) How many moles of gas are needed to produce 100 moles of  $\text{C}_2\text{H}_5\text{O}^+$  and 100 moles of  $\text{C}_2\text{H}_5^+$ ?

$$\text{C}_2\text{H}_6\text{O} + \text{H}^+ \rightarrow \text{C}_2\text{H}_7\text{O}^+ \rightarrow \text{C}_2\text{H}_5\text{O}^+ + \text{H}_2 \rightarrow \text{C}_2\text{H}_5^+ + \text{H}_2\text{O}$$

2. Molecular formulae:  $\text{C}_2\text{H}_6\text{O} + \text{H}^+ \rightarrow \text{C}_2\text{H}_7\text{O}^+ \rightarrow \text{C}_2\text{H}_5\text{O}^+ + \text{H}_2 \rightarrow \text{C}_2\text{H}_5^+ + \text{H}_2\text{O}$

(A) How many moles of hydrogen molecules are required to produce 100 moles of  $\text{C}_2\text{H}_7\text{O}^+$ ?

$$\text{C}_2\text{H}_6\text{O} + \text{H}^+ \rightarrow \text{C}_2\text{H}_7\text{O}^+ \rightarrow \text{C}_2\text{H}_5\text{O}^+ + \text{H}_2 \rightarrow \text{C}_2\text{H}_5^+ + \text{H}_2\text{O}$$

(B) How many moles of hydrogen molecules are required to produce 100 moles of  $\text{C}_2\text{H}_5\text{O}^+$ ?

$$\text{C}_2\text{H}_6\text{O} + \text{H}^+ \rightarrow \text{C}_2\text{H}_7\text{O}^+ \rightarrow \text{C}_2\text{H}_5\text{O}^+ + \text{H}_2 \rightarrow \text{C}_2\text{H}_5^+ + \text{H}_2\text{O}$$

(C) How many moles of water are produced when 100 moles of  $\text{C}_2\text{H}_7\text{O}^+$  are made?

$$\text{C}_2\text{H}_6\text{O} + \text{H}^+ \rightarrow \text{C}_2\text{H}_7\text{O}^+ \rightarrow \text{C}_2\text{H}_5\text{O}^+ + \text{H}_2 \rightarrow \text{C}_2\text{H}_5^+ + \text{H}_2\text{O}$$

3. Calculate the mass of hydrogen gas produced when 1.00 moles of ethanol reacts with oxygen



$$\text{C}_2\text{H}_5\text{OH} + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$$

Answer: (A) 1.00 g (B) 1.50 g (C) 2.00 g (D) 3.00 g

# Chemistry Stoichiometry Problem Sheet 2 Answer Key

**CO Houle**



## **Chemistry Stoichiometry Problem Sheet 2 Answer Key:**

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### **Table of Contents Chemistry Stoichiometry Problem Sheet 2 Answer Key**

1. Understanding the eBook Chemistry Stoichiometry Problem Sheet 2 Answer Key
  - The Rise of Digital Reading Chemistry Stoichiometry Problem Sheet 2 Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Stoichiometry Problem Sheet 2 Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Stoichiometry Problem Sheet 2 Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Stoichiometry Problem Sheet 2 Answer Key

- Personalized Recommendations
  - Chemistry Stoichiometry Problem Sheet 2 Answer Key User Reviews and Ratings
  - Chemistry Stoichiometry Problem Sheet 2 Answer Key and Bestseller Lists
5. Accessing Chemistry Stoichiometry Problem Sheet 2 Answer Key Free and Paid eBooks
    - Chemistry Stoichiometry Problem Sheet 2 Answer Key Public Domain eBooks
    - Chemistry Stoichiometry Problem Sheet 2 Answer Key eBook Subscription Services
    - Chemistry Stoichiometry Problem Sheet 2 Answer Key Budget-Friendly Options
  6. Navigating Chemistry Stoichiometry Problem Sheet 2 Answer Key eBook Formats
    - ePub, PDF, MOBI, and More
    - Chemistry Stoichiometry Problem Sheet 2 Answer Key Compatibility with Devices
    - Chemistry Stoichiometry Problem Sheet 2 Answer Key Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chemistry Stoichiometry Problem Sheet 2 Answer Key
    - Highlighting and Note-Taking Chemistry Stoichiometry Problem Sheet 2 Answer Key
    - Interactive Elements Chemistry Stoichiometry Problem Sheet 2 Answer Key
  8. Staying Engaged with Chemistry Stoichiometry Problem Sheet 2 Answer Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chemistry Stoichiometry Problem Sheet 2 Answer Key
  9. Balancing eBooks and Physical Books Chemistry Stoichiometry Problem Sheet 2 Answer Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chemistry Stoichiometry Problem Sheet 2 Answer Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chemistry Stoichiometry Problem Sheet 2 Answer Key
    - Setting Reading Goals Chemistry Stoichiometry Problem Sheet 2 Answer Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chemistry Stoichiometry Problem Sheet 2 Answer Key

- Fact-Checking eBook Content of Chemistry Stoichiometry Problem Sheet 2 Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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