

Introduction

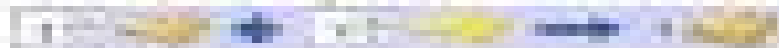
This course is a general introduction to the field of chemistry. It is designed to give students a solid foundation in the basic principles of chemistry, including atomic structure, chemical bonding, and stoichiometry. The course is intended for students who are interested in pursuing a degree in chemistry or a related field. The course is divided into two semesters. The first semester covers the basic principles of chemistry, and the second semester covers the more advanced topics of chemistry.

Prerequisites: High school chemistry or equivalent. [View the Prerequisite Chart for the Chemistry Program](#) for more information.

Part 1: Atomic Structure



1. Define the terms atom, element, and compound. Explain the difference between an atom and a molecule. Describe the structure of an atom, including the nucleus and the electron cloud.
2. Explain the periodic table of elements. Describe the properties of the elements in each group and period.
3. Calculate the atomic mass of an element from its isotopic composition.



4. Calculate the molar mass of a compound from its chemical formula.
5. Calculate the number of moles of a substance from its mass and molar mass.

Element	Atomic Mass	Molar Mass	Number of Atoms	Mass of Sample	Number of Moles
H ₂	1.008 g/mol	2.016 g/mol	1.000 mol	2.016 g	0.500 mol
O ₂	16.00 g/mol	32.00 g/mol	0.500 mol	16.00 g	0.156 mol
CO ₂	44.01 g/mol	44.01 g/mol	0.500 mol	22.005 g	0.113 mol
NaCl	58.44 g/mol	58.44 g/mol	0.500 mol	29.22 g	0.085 mol

6. Calculate the percent composition of a compound from its chemical formula.
7. Calculate the empirical formula of a compound from its percent composition.

Part 2: Chemical Bonding



8. Explain the difference between ionic and covalent bonding. Describe the properties of ionic and covalent compounds.
9. Calculate the lattice energy of an ionic compound from its enthalpy of formation.
10. Calculate the bond energy of a covalent bond from its enthalpy of formation.

$$\Delta H_{\text{formation}} = \sum \Delta H_{\text{bond breaking}} - \sum \Delta H_{\text{bond forming}}$$

11. Calculate the enthalpy of formation of a compound from its enthalpy of combustion.

Compound	Enthalpy of Formation (kJ/mol)	Enthalpy of Combustion (kJ/mol)	Enthalpy of Formation (kJ/mol)	Enthalpy of Formation (kJ/mol)
H ₂ O(l)	-285.8	-285.8	-285.8	-285.8
CO ₂ (g)	-393.5	-393.5	-393.5	-393.5
CH ₄ (g)	-74.8	-890.3	-74.8	-74.8
C ₂ H ₆ (g)	-84.7	-1560.7	-84.7	-84.7

Work on these problems. Show your work on the next page.

Basic Stoichiometry Phet Lab Homework Exercises Answers

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