

## CHAPTER 9 REVIEW

**Stoichiometry****SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

- \_\_\_\_\_ The coefficients in a chemical equation represent the
  - masses in grams of all reactants and products.
  - relative number of moles of reactants and products.
  - number of atoms of each element in each compound in a reaction.
  - number of valence electrons involved in a reaction.
- \_\_\_\_\_ Which of the following would not be studied within the topic of stoichiometry?
  - the mole ratio of Al to Cl in the compound aluminum chloride
  - the mass of carbon produced when a known mass of sucrose decomposes
  - the number of moles of hydrogen that will react with a known quantity of oxygen
  - the amount of energy required to break the ionic bonds in  $\text{CaF}_2$
- \_\_\_\_\_ A balanced chemical equation allows you to determine the
  - mole ratio of any two substances in the reaction.
  - energy released in the reaction.
  - electron configuration of all elements in the reaction.
  - reaction mechanism involved in the reaction.
- \_\_\_\_\_ The relative number of moles of hydrogen to moles of oxygen that react to form water represents a(n)
  - reaction sequence.
  - bond energy.
  - mole ratio.
  - element proportion.
- Given the reaction represented by the following unbalanced equation:  
 $\text{N}_2\text{O}(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{NO}_2(\text{g})$ 
  - Balance the equation.

\_\_\_\_\_

b. What is the mole ratio of  $\text{NO}_2$  to  $\text{O}_2$ ?

\_\_\_\_\_

c. If 20.0 mol of  $\text{NO}_2$  form, how many moles of  $\text{O}_2$  must have been consumed?

\_\_\_\_\_

d. Twice as many moles of  $\text{NO}_2$  form as moles of  $\text{N}_2\text{O}$  are consumed. True or False?

\_\_\_\_\_

e. Twice as many grams of  $\text{NO}_2$  form as grams of  $\text{N}_2\text{O}$  are consumed. True or False?

# Chapter 9 Stoichiometry Answers

**S Baum**



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**The Practice of Chemistry Study Guide & Solutions Manual** Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

**Chemistry Workbook For Dummies** Peter J. Mikulecky, Katherine Brutlag, Michelle Rose Gilman, Brian Peterson, 2008-08-06 From liquids and solids to acids and bases work chemistry equations and use formulas with ease Got a grasp on the chemistry terms and concepts you need to know but get lost halfway through a problem or worse yet not know where to begin Have no fear this hands on guide helps you solve many types of chemistry problems in a focused step by step manner With problem solving shortcuts and lots of practice exercises you ll build your chemistry skills and improve your performance both in and out of the science lab You ll see how to work with numbers atoms and elements make and remake compounds understand changes in terms of energy make sense of organic chemistry and more 100s of Problems Know where to begin and how to solve the most common chemistry problems Step by step answer sets clearly identify where you went wrong or right with a problem Understand the key exceptions to chemistry rules Use chemistry in practical applications with confidence

High School Chemistry Michael A. Buben, 2018-05-18 The Teacher s Edition is designed to add direction to High School Chemistry which is an outline of notes The books go hand in hand Modified Second Edition as a student s workbook and guidance from the Teacher s Edition

**Scientific Soapmaking** Kevin M. Dunn, 2010 Scientific Soapmaking bridges the gap between the technical and craft literature It explains the chemistry of fats oils and soaps and teaches sophisticated analytical techniques that can be carried out using equipment and materials familiar to makers of handcrafted soap

**Answer Book to Accompany Boikess/Edelson, Chemical Principles, Third Edition** Robert S. Boikess, 1985

Student Solutions Manual, Chemistry, Principles & Practice, Second Edition John M. DeKorte, 1997

**Programmed Problem-solving for First-year Chemistry** Arnold B. Loebel, 1983

**Instructor's Manual and Test Bank to Accompany Basic Concepts of Chemistry** Malone, 2000-04-18

**Solutions to Red Exercises - Chemistry** Roxy Wilson, 2005-05 Prepared by Roxy Wilson of University of Illinois Urbana Champaign Full so lutions to all of the red numbered exercises in the text are pro vided Short answers to red exercises are found in the appendix of the text

Solutions to Black Exercises Roxy Wilson, 1999-08

Solutions Manual to Accompany Chemistry Raymond Chang, 1984

**Basic Concepts of Chemistry** Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating

to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance

**Solutions Manual to Accompany Chemical Principles** Robert Boikess,1994 *Student Solutions Manual to Accompany Chemistry & Chemical Reactivity, Fourth Edition, Kotz & Treichel* Alton J. Banks,Paul Treichel,1999

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