

Name _____ Date _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then use data for table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of O_2 , H_2 , and H_2O are 32.00 g/mol, 2.016 g/mol, and 18.02 g/mol, respectively.

Amount of O_2	Amount of H_2	Amount of H_2O	Limiting Reactant	Amount of Excess Reactant
1 molecule	1 molecule	1 molecule	none	none
2 molecules	2 molecules	2 molecules	H_2	1 molecule O_2
2 molecules	3 molecules	3 molecules	O_2	1 molecule H_2
1 mol	2 mol	2 mol	H_2	1 mol O_2
2 mol	3 mol	3 mol	O_2	1 mol H_2
1 mol	3 mol	2 mol	H_2	1 mol O_2
1 mol	4 mol	2 mol	H_2	2 mol O_2
1 mol	5 mol	2 mol	H_2	3 mol O_2
1 mol	6 mol	2 mol	H_2	4 mol O_2
1 mol	7 mol	2 mol	H_2	5 mol O_2
1 mol	8 mol	2 mol	H_2	6 mol O_2
1 mol	9 mol	2 mol	H_2	7 mol O_2
1 mol	10 mol	2 mol	H_2	8 mol O_2
1 mol	11 mol	2 mol	H_2	9 mol O_2
1 mol	12 mol	2 mol	H_2	10 mol O_2
1 mol	13 mol	2 mol	H_2	11 mol O_2
1 mol	14 mol	2 mol	H_2	12 mol O_2
1 mol	15 mol	2 mol	H_2	13 mol O_2
1 mol	16 mol	2 mol	H_2	14 mol O_2
1 mol	17 mol	2 mol	H_2	15 mol O_2
1 mol	18 mol	2 mol	H_2	16 mol O_2
1 mol	19 mol	2 mol	H_2	17 mol O_2
1 mol	20 mol	2 mol	H_2	18 mol O_2

balanced equation mole ratio = 2 mol H_2 /1 mol O_2

32.00 g O_2 = 1 mol O_2 /32.00 g O_2 = 0.3125 mol O_2

20.00 g H_2 = 1 mol H_2 /2.016 g H_2 = 9.920 mol H_2

actual mole ratio = 9.920 mol H_2 /0.3125 mol O_2 = 31.44 mol H_2 /1 mol O_2

Because the actual mole ratio of H_2 / O_2 is larger than the balanced equation mole ratio of H_2 / O_2 , there is an excess of H_2 ; O_2 is the limiting reactant.

Mass of H_2 used = 0.3125 mol O_2 \times 2 mol H_2 /1 mol O_2 = 0.625 mol H_2

0.625 mol H_2 \times 2.016 g H_2 /1 mol H_2 = 1.26 g H_2

Mass of H_2O produced = 0.625 mol H_2 \times 1 mol H_2O /1 mol H_2 = 0.625 mol H_2O

0.625 mol H_2O \times 18.02 g H_2O /1 mol H_2O = 11.26 g H_2O

Excess H_2 = 20.00 g H_2 - 1.26 g H_2 = 18.74 g H_2

Excess O_2 = 32.00 g O_2 - 32.00 g O_2 = 0 g O_2

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CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of gallium nitride, a substance used in the manufacturing of some semiconductors for lasers.



In one experiment, the reaction yielded 7.42 g of the oxide from a 1.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_3N_2 are 69.72 g/mol and 177.48 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

1. actual yield volume = 7.42 g Ga_3N_2
2. mass of reactant = 1.00 g Ga
3. number of moles of reactant = 1.00 g Ga \times $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}}$ = 1.43 mol Ga
4. number of moles of product = 0.95 mol Ga \times $\frac{2 \text{ mol Ga}_3\text{N}_2}{3 \text{ mol Ga}}$ = 0.63 mol Ga_3N_2
5. theoretical yield = 0.0000 mol Ga_3N_2 = $\frac{177.48 \text{ g Ga}_3\text{N}_2}{1 \text{ mol Ga}_3\text{N}_2}$ correct
6. percent yield = $\frac{7.42 \text{ g Ga}_3\text{N}_2}{177.48 \text{ g Ga}_3\text{N}_2} \times 100$ = 4.18% correct

Study Guide For Content Mastery Answer Key Chapter 12

Margaret Ann Richek



Study Guide For Content Mastery Answer Key Chapter 12:

Study Guide For Content Mastery Answer Key Chapter 12 Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Study Guide For Content Mastery Answer Key Chapter 12**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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