



# COVALENT BONDING

## Practice Problems

Do your answers to the following problems.

### SECTION 5.1 MOLECULAR COMPOUNDS

- Draw the following molecules using Lewis dot notation.
  - $\text{H}_2$  (hydrogen)
  - $\text{O}_2$  (oxygen)
  - $\text{N}_2$  (nitrogen)
  - $\text{H}_2\text{O}$  (water)
  - $\text{CO}_2$  (carbon dioxide)
  - $\text{CH}_4$  (methane)
  - $\text{NH}_3$  (ammonia)
  - $\text{H}_2\text{S}$  (hydrogen sulfide)
  - $\text{C}_2\text{H}_6$  (ethane)
  - $\text{C}_2\text{H}_4$  (ethene)
  - $\text{C}_2\text{H}_2$  (ethyne)
- Draw the following molecules using Lewis dot notation.
  - $\text{H}_2\text{O}_2$  (hydrogen peroxide)
  - $\text{C}_2\text{H}_5\text{OH}$  (ethanol)
  - $\text{C}_2\text{H}_5\text{NH}_2$  (ethanamine)
  - $\text{C}_2\text{H}_5\text{SH}$  (ethanethiol)
  - $\text{C}_2\text{H}_5\text{Cl}$  (chloroethane)
  - $\text{C}_2\text{H}_5\text{Br}$  (bromoethane)
  - $\text{C}_2\text{H}_5\text{I}$  (iodoethane)
  - $\text{C}_2\text{H}_5\text{F}$  (fluoroethane)
  - $\text{C}_2\text{H}_5\text{NO}_2$  (nitroethane)
  - $\text{C}_2\text{H}_5\text{SO}_2\text{CH}_3$  (methyl sulfonate)
  - $\text{C}_2\text{H}_5\text{SO}_2\text{NH}_2$  (sulfonamide)
- Draw the following molecules using Lewis dot notation.
  - $\text{C}_2\text{H}_5\text{COOH}$  (propanoic acid)
  - $\text{C}_2\text{H}_5\text{COCH}_3$  (propanone)
  - $\text{C}_2\text{H}_5\text{CONH}_2$  (ethanamide)
  - $\text{C}_2\text{H}_5\text{COCl}$  (propanoyl chloride)
  - $\text{C}_2\text{H}_5\text{COBr}$  (propanoyl bromide)
  - $\text{C}_2\text{H}_5\text{COI}$  (propanoyl iodide)
  - $\text{C}_2\text{H}_5\text{COF}$  (propanoyl fluoride)
  - $\text{C}_2\text{H}_5\text{CO}_2\text{CH}_3$  (methyl propanoate)
  - $\text{C}_2\text{H}_5\text{CO}_2\text{NH}_2$  (propanamide)
  - $\text{C}_2\text{H}_5\text{CO}_2\text{NHC}_2\text{H}_5$  (diethyl propanoate)
  - $\text{C}_2\text{H}_5\text{CO}_2\text{C}_2\text{H}_5$  (diethyl propanoate)

### SECTION 5.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis dot structure for  $\text{H-F}$ .
- Draw the Lewis dot structure for  $\text{H}_2\text{O}$ .
- Draw the Lewis dot structure for  $\text{CH}_4$ .
- Draw the Lewis dot structure for  $\text{C}_2\text{H}_6$ .
- Draw the Lewis dot structure for  $\text{C}_2\text{H}_4$ .
- Draw the Lewis dot structure for  $\text{C}_2\text{H}_2$ .
- Draw the Lewis dot structure for  $\text{H-C}\equiv\text{C-H}$ .
- Draw the Lewis dot structure for  $\text{C}_2\text{H}_5\text{OH}$ .



### SECTION 5.3 EXERCISES: THE LEXIS

- Write the Lewis dot structure for  $\text{C}_2\text{H}_6$ .
- Write the Lewis dot structure for  $\text{C}_2\text{H}_4$ .
- Write the Lewis dot structure for  $\text{C}_2\text{H}_2$ .
- Write the Lewis dot structure for  $\text{C}_2\text{H}_5\text{OH}$ .
- Write the Lewis dot structure for  $\text{C}_2\text{H}_5\text{COOH}$ .
- Write the Lewis dot structure for  $\text{C}_2\text{H}_5\text{COCH}_3$ .
- Write the Lewis dot structure for  $\text{C}_2\text{H}_5\text{CONH}_2$ .
- Write the Lewis dot structure for  $\text{C}_2\text{H}_5\text{COCl}$ .
- Write the Lewis dot structure for  $\text{C}_2\text{H}_5\text{COBr}$ .
- Write the Lewis dot structure for  $\text{C}_2\text{H}_5\text{COI}$ .
- Write the Lewis dot structure for  $\text{C}_2\text{H}_5\text{COF}$ .
- Write the Lewis dot structure for  $\text{C}_2\text{H}_5\text{CO}_2\text{CH}_3$ .
- Write the Lewis dot structure for  $\text{C}_2\text{H}_5\text{CO}_2\text{NH}_2$ .
- Write the Lewis dot structure for  $\text{C}_2\text{H}_5\text{CO}_2\text{NHC}_2\text{H}_5$ .
- Write the Lewis dot structure for  $\text{C}_2\text{H}_5\text{CO}_2\text{C}_2\text{H}_5$ .

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**Clifford Lane**



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