

20 Points

Name: Phyllis A. Smith Lab Section: 200

Suppose that you had the correct equation for the following combustion reaction. What highest boiling point compound?

$\text{C}_2\text{H}_6\text{O}$

C_2H_6

C_2H_4

C_2H_2

$\text{C}_2\text{H}_6\text{O}$
Ethanol

Diethyl ether

C_2H_4

C_2H_2
Acetylene

1. If you have the correct equation for the combustion of C_2H_6 , how many moles of H_2O would be produced by combustion of 100 g of C_2H_6 ?

$$\frac{100 \text{ g } \text{C}_2\text{H}_6}{30 \text{ g}} \left| \frac{1 \text{ mol}}{30 \text{ g}} \right| \times \frac{7}{1} \left| \frac{1 \text{ mol}}{1 \text{ mol}} \right| = 2.33 \text{ mol } \text{H}_2\text{O}$$

2.33 mol

2. If you were in the market and wanted the most expensive of C_2H_6 , C_2H_4 or C_2H_2 based on the density of the solution in C_2H_6 ?

$$\text{C}_2\text{H}_6: \frac{10 \text{ g } \text{C}_2\text{H}_6}{100 \text{ g } \text{soln}} \left| \frac{1 \text{ mol}}{30 \text{ g}} \right| \left| \frac{1 \text{ mol}}{1 \text{ mol}} \right| = 0.33 \text{ mol } \text{C}_2\text{H}_6$$

$$\text{C}_2\text{H}_4: \frac{10 \text{ g } \text{C}_2\text{H}_4}{100 \text{ g } \text{H}_2\text{O}} \left| \frac{1 \text{ mol}}{28 \text{ g}} \right| \left| \frac{1 \text{ mol}}{1 \text{ mol}} \right| = 0.36 \text{ mol } \text{C}_2\text{H}_4$$

0.36 mol

Acetylene

0.37 mol

Acetylene

Chem 110 Lab Questions And Answers

Lyman Churchill Newell



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