

Periodic Table Trend Activities:



Purpose: The Periodic Table is a useful tool that helps us predict properties of various elements. In this activity, we will look at 3 properties:

1. **Atomic radius:** The distance from the atomic nucleus to the outmost occupied orbital in a atom
2. **Electronegativity:** The ability of an atom to attract a bonded pair of electrons from another atom
3. **Ionization Energy:** The energy needed to remove an electron from a neutral gaseous atom
4. **Electron Affinity:** The energy released when an electron is added to a neutral gaseous atom

Procedure: Click on the [Cool Periodic Table link](#) (click on the element symbol). Record the atomic radius, electronegativity, and first ionization energy for each element in the table below:

<http://www.lyncsburg.net/hhs/chemistry/trends/>

Data:

Element	Atomic Radius	Electronegativity	1 st Ionization Energy
Li			
Be			
B			
Na			
Mg			
Al			
C			
Si			
N			
P			

Discussion:

1. What appears to be the trend in atomic radius as you move from left to right in a row?
2. What appears to be the trend in atomic radius as you move down a column?

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Table of Contents Chemistry Lab Periodic Trends Computer Activity Answers

1. Understanding the eBook Chemistry Lab Periodic Trends Computer Activity Answers
 - The Rise of Digital Reading Chemistry Lab Periodic Trends Computer Activity Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Lab Periodic Trends Computer Activity Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Lab Periodic Trends Computer Activity Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Lab Periodic Trends Computer Activity Answers
 - Personalized Recommendations
 - Chemistry Lab Periodic Trends Computer Activity Answers User Reviews and Ratings
 - Chemistry Lab Periodic Trends Computer Activity Answers and Bestseller Lists
5. Accessing Chemistry Lab Periodic Trends Computer Activity Answers Free and Paid eBooks
 - Chemistry Lab Periodic Trends Computer Activity Answers Public Domain eBooks
 - Chemistry Lab Periodic Trends Computer Activity Answers eBook Subscription Services
 - Chemistry Lab Periodic Trends Computer Activity Answers Budget-Friendly Options

6. Navigating Chemistry Lab Periodic Trends Computer Activity Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Lab Periodic Trends Computer Activity Answers Compatibility with Devices
 - Chemistry Lab Periodic Trends Computer Activity Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Lab Periodic Trends Computer Activity Answers
 - Highlighting and Note-Taking Chemistry Lab Periodic Trends Computer Activity Answers
 - Interactive Elements Chemistry Lab Periodic Trends Computer Activity Answers
8. Staying Engaged with Chemistry Lab Periodic Trends Computer Activity Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Lab Periodic Trends Computer Activity Answers
9. Balancing eBooks and Physical Books Chemistry Lab Periodic Trends Computer Activity Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Lab Periodic Trends Computer Activity Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Lab Periodic Trends Computer Activity Answers
 - Setting Reading Goals Chemistry Lab Periodic Trends Computer Activity Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Lab Periodic Trends Computer Activity Answers
 - Fact-Checking eBook Content of Chemistry Lab Periodic Trends Computer Activity Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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