

# QUESTION 1

Volume of pipette used  $25.00 \text{ cm}^3$   
 (V<sub>0</sub>)

Titration	Rough	1	2	3
Final burette reading (cm <sup>3</sup> )	20.80	41.40	20.40	40.90
Initial burette reading (cm <sup>3</sup> )	0.00	20.80	0.00	20.40
Volume of A used	20.80	20.60	20.40	20.50

② <sup>average</sup> Volume of A used =  $\frac{20.60 + 20.40 + 20.50}{3}$   
 $= 20.50 \text{ cm}^3$

③(i) Concentration of A in mol/dm<sup>3</sup> =  $\frac{\text{conc. in g/dm}^3}{\text{molar mass}}$   
 $= \frac{4.4 \text{ g/dm}^3}{2[13] + 32 + 4[16]}$   
 $= \frac{4.4}{98} = 0.0449 \text{ mol/dm}^3$

③(ii) Concentration of B in mol/dm<sup>3</sup>, C<sub>B</sub>

Using  $\frac{C_A V_A}{C_B V_B} = \frac{n_A}{n_B}$

$$\frac{0.0449 \times 20.50}{C_B \times 25} = \frac{1}{1}$$

$$C_B = \frac{0.0449 \times 20.50}{25} = 0.03287 \text{ mol/dm}^3$$

# Chemistry Practical Question And Answer

**Wolfgang Guggemos**



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