

- (c) 350 cm, convergent
(d) 150 cm, divergent
16. Three resistances A, B and C each of 5Ω are connected to a battery of 10 V as shown in the figure. Calculate the current through C:
-
- (a) 2 amp
(b) 1 amp
(c) 3 amp
(d) 0.6 amp
17. The device used to generate electrical energy is:
- (a) Electric motor
(b) Generator
(c) Galvanometer
(d) Voltmeter
18. The number of molecules in 11 g of CO_2 are:
- (a) 0.25×10^{23}
(b) 0.50×10^{23}
(c) 1.00×10^{23}
(d) 1.51×10^{23}
19. The value of charge / mass ratio of electron was determined by:
- (a) W.K. Roentgen
(b) J.J. Thomson
(c) Marie Curie
(d) Niels Bohr
20. Which of the following reactions is an example of combination reaction?
- (a) $\text{Zn}(s) + \text{CuSO}_4(aq) \rightarrow \text{ZnSO}_4(aq) + \text{Cu}(s)$
(b) $\text{BaCl}_2(aq) + \text{Na}_2\text{SO}_4(aq) \rightarrow \text{BaSO}_4(s) + 2\text{NaCl}(aq)$
(c) $\text{CaCO}_3(s) \rightarrow \text{CaO}(s) + \text{CO}_2(g)$
(d) $\text{C}(s) + \text{O}_2(g) \rightarrow \text{CO}_2(g)$
21. The element 'X' and 'Y' have atomic numbers 12 and 17 respectively. Element 'X' reacts with element 'Y' to form compound with molecular formula:
- (a) XY
(b) XY_2
(c) XY_3
(d) X_3Y_2
22. The atomic radius (pm) of Li, Na, K and Rb varies in the order:
- (a) $\text{Na} < \text{K} < \text{Rb} < \text{Li}$
(b) $\text{K} < \text{Na} < \text{Li} < \text{Rb}$
(c) $\text{Li} < \text{Na} < \text{K} < \text{Rb}$
(d) $\text{Rb} < \text{K} < \text{Na} < \text{Li}$
23. The electron affinity value (kJ mol^{-1}) of fluorine (F) is less than:
- (a) Hydrogen (H)
(b) Lithium (Li)
(c) Oxygen (O)
(d) Chlorine (Cl)
24. Which of the following is an example of strong electrolyte?
- (a) H_2CO_3
(b) NH_4OH
(c) NaCl
(d) $\text{COOH} - \text{COOH}$
25. If pH of a solution changes from 5 to 4, the change in hydrogen ion concentration shall be:
- (a) two times
(b) five times
(c) ten times
(d) twenty times
26. Bleaching powder is manufactured by the reaction of:
- (a) CaCl_2 and CaCO_3
(b) Cl_2 and $\text{Ca}(\text{OH})_2$
(c) Cl_2 and CaSO_4
(d) Cl_2 and $\text{Ca}(\text{HCO}_3)_2$
- The hard glass is obtained by fusing:
- (a) soda ash, sand and limestone
(b) a mixture of sand, lime, borax and alkali carbonates
(c) potassium carbonate and limestone
(d) potassium carbonate, lead oxide and sand
28. Which of the following metals is most reactive?
- (a) Aluminium
(b) lead
(c) Mercury
(d) Silver
29. Which of the following alloys contains chromium?
- (a) Steel
(b) Stainless steel
(c) Magnalium
(d) Brass
30. Which of the following is monomer of natural rubber?
- (a) Chloroethene
(b) Chloroprene
(c) Isoprene
(d) Buta-1, 3-diene
31. Alkaline KMnO_4 oxidises propanone to:
- (a) propanoic acid
(b) ethanoic acid
(c) methanoic acid
(d) oxalic acid
32. Which of the following metals can displace Zn from ZnSO_4 solution?
- (a) Calcium
(b) Copper
(c) Iron
(d) Mercury
33. Heating of sodium ethanoate with soda lime yields:
- (a) ethane
(b) methane
(c) ethanol
(d) methanol
34. The compound formed by the reaction of ethyne with bromine is:
- (a) $\text{Br}-\text{CH}=\text{CH}-\text{Br}$
(b) $\text{Br}-\text{CH}_2-\text{CHBr}_2$
(c) $\text{CH}_2=\text{CH}-\text{Br}$
(d) $\text{Br}_3\text{CH}-\text{CHBr}_2$
35. Cell were first discovered by Robert Hooke in the year:
- (a) 1665
(b) 1674
(c) 1831
(d) 1839
36. The plant cells have a rigid cell wall that lies:
- (a) Outside the plasma membrane
(b) Inside the plasma membrane
(c) In between the plasma membranes
(d) None of the above
37. During mitosis nucleolus and nuclear membrane are lost in stage:
- (a) Prophase
(b) Metaphase
(c) Anaphase
(d) Telophase
38. An example of simple and permanent tissue is:
- (a) Xylem
(b) Phloem
(c) Sclerenchyma
(d) All the above
39. The shape of squamous epithelial tissue is:
- (a) Cubical
(b) Flattened
(c) Pillar like
(d) None of the above
40. Fresh water sponge belongs to phylum:
- (a) Arthropoda
(b) Annelida

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Table of Contents Rpvv Entrance Question Paper

1. Understanding the eBook Rpvv Entrance Question Paper
 - The Rise of Digital Reading Rpvv Entrance Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Rpvv Entrance Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Rpvv Entrance Question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Rpvv Entrance Question Paper

- Personalized Recommendations
 - Rpvv Entrance Question Paper User Reviews and Ratings
 - Rpvv Entrance Question Paper and Bestseller Lists
5. Accessing Rpvv Entrance Question Paper Free and Paid eBooks
 - Rpvv Entrance Question Paper Public Domain eBooks
 - Rpvv Entrance Question Paper eBook Subscription Services
 - Rpvv Entrance Question Paper Budget-Friendly Options
 6. Navigating Rpvv Entrance Question Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Rpvv Entrance Question Paper Compatibility with Devices
 - Rpvv Entrance Question Paper Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Rpvv Entrance Question Paper
 - Highlighting and Note-Taking Rpvv Entrance Question Paper
 - Interactive Elements Rpvv Entrance Question Paper
 8. Staying Engaged with Rpvv Entrance Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Rpvv Entrance Question Paper
 9. Balancing eBooks and Physical Books Rpvv Entrance Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Rpvv Entrance Question Paper
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Rpvv Entrance Question Paper
 - Setting Reading Goals Rpvv Entrance Question Paper
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Rpvv Entrance Question Paper

- Fact-Checking eBook Content of Rpvv Entrance Question Paper
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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