

Chemistry I (H) - Unit 02 (Pt. II - Nuclear Chemistry) Study Guide

- 1) Define, compare, and contrast radiation (pg. 122) to radioactivity (pg. 122).
- 2) Define AND give an example of a radioisotope (pg. 861).
- 3) Define nuclear reactions.
 - Define nuclear equations (pg. 123/869) -
 - Example:
- 4) Define and provide examples of the following forms of radioactive decay:
 - alpha decay (pg. 867) -
 - alpha particle (pg. 123/862) -
 - Example Nuclear Equation:
 - beta decay (pg. 867) -
 - beta particle (pg. 123/863) -
 - Example Nuclear Equation:
 - gamma radiation (pg. 124) -
 - gamma ray (pg. 124/863) -
 - Example Nuclear Equation:
 - Study the Table 24.2 from pg. 861 that we drew in our notes about the properties of decay (penetrating power, mass, etc.)
- 5) Define AND give an example of transmutation (pg. 865).
- 6) What do we mean by the words emit or emanate? Give an example.
- 7) Compare a *stable nucleus* to that of an *unstable nucleus*.

Section 25 Nuclear Chemistry Study Guide Answers

Robert K. Wismer



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Nuclear Science Abstracts, 1971

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General Chemistry Robert K. Wismer, 1993

Fundamentals of Chemistry Fred H. Redmore, 1979

Chemistry George M. Bodner, Harry L. Pardue, 1989-01-23 A modern experimental approach to first year chemistry This unique introductory account employs experimental observations to construct the principles of general chemistry An early introduction to observable descriptive chemistry lays the basis for the well developed exposition that follows

Chemistry and Life John William Hill, Dorothy M. Feigl, Stuart J. Baum, 1993

Organic Chemistry Stanislaw Skonieczny, George H. Schmid, 1996

General Chemistry with Qualitative Analysis Ralph H. Petrucci, Robert K. Wismer, 1987

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