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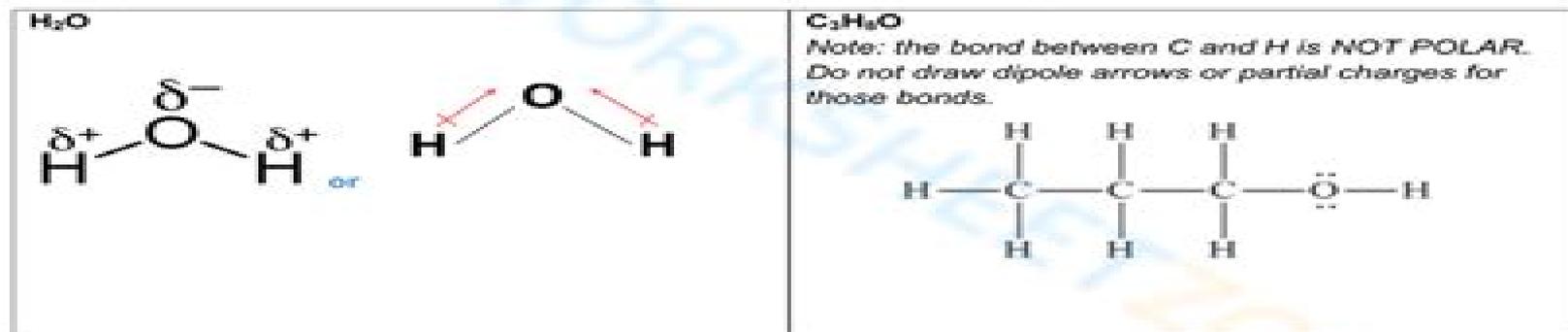
Intermolecular Forces Lab Worksheet - Answer Key



Directions: Today you will complete two experiments that compare the properties of water to isopropanol (a type of rubbing alcohol). Before conducting each experiment, you will first read about the experiment and then make a prediction about what you think will happen. Afterward you do the experiment, you will write an explanation of what was happening at the atomic level that let us observe these properties.

Introduction: In many ways, water is a miracle liquid. It is essential for all living things on earth and it is often referred to as a universal solvent because many substances dissolve in it. Water displays unusual properties due to the ways the individual water molecules interact with each other.

1. Draw the structural formula (with lone pairs) for the two molecules we will be comparing today, water (H_2O) and isopropanol ($\text{C}_3\text{H}_8\text{O}$). Draw a dipole arrow to show the direction the electrons are being pulled and label the areas of partial positive and negative charges.



2. Which molecule is more polar? water

Experiment #1: Penny

How many drops of water fit on the head of a penny? How many drops of isopropanol fit on the head of a penny?

Prediction:

- a. I think _____ drops of water will fit on the head of the penny because _____

Intermolecular Forces Lab And Answers

Alan Hinchliffe



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Lab Manual for General, Organic, and Biochemistry Denise Guinn, Rebecca Brewer, 2009-08-21 Teaching all of the necessary concepts within the constraints of a one term chemistry course can be challenging Authors Denise Guinn and Rebecca Brewer have drawn on their 14 years of experience with the one term course to write a textbook that incorporates biochemistry and organic chemistry throughout each chapter emphasizes cases related to allied health and provides students with the practical quantitative skills they will need in their professional lives Essentials of General Organic and Biochemistry captures student interest from day one with a focus on attention getting applications relevant to health care professionals and as much pertinent chemistry as is reasonably possible in a one term course Students value their experience with chemistry getting a true sense of just how relevant it is to their chosen profession To browse a sample chapter view sample ChemCasts and more visit www.whfreeman.com gob

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Lab Manual T/a Human Physiology Bill W. Tillery, 2001-03

Take-Home Chemistry Michael Horton, 2011 For high school science teachers homeschoolers science coordinators and informal science educators this collection of 50 inquiry based labs provides hands on ways for students to learn science at home safely Author Michael Horton promises that students who conduct the labs in Take Home Chemistry as supplements to classroom instruction will enhance higher level thinking improve process skills and raise high stakes test scores

Air Force Research

Resumés, 1961 **Phy. Lab and Pocket Lab Wk/Sheets Phy:P&P** Zitzewitz, 1998-07 Holt Chemistry Ralph Thomas Myers, 2004 Air Force Research Resumés, 1964 **McGraw-Hill's 500 MCAT Organic Chemistry Questions to Know by Test Day** John T. Moore, Richard H. Langley, 2012-06-19 Contains five hundred questions and answers about organic chemistry to help students prepare for the MCAT exam *Exploring Chemistry (Loose-Leaf)* Matthew Johll, 2012-05-04 Matthew Johll's *Exploring Chemistry* covers the standard topics for the nonmajors course in the typical order but each chapter unfolds in the context of a single case study that helps students connect what they are learning to real life situations For example students work through the often difficult topics of molecular structure gas laws and organic chemistry by learning about the development of powerful new chemotherapy drugs new technologies for screening airline passengers and the creation of biodegradable biopolymers It's the same case driven approach that Johll uses in his acclaimed *Investigating Chemistry* now in its Third Edition but *Exploring Chemistry* goes beyond the other book's specific focus on examples from forensic science to use real life stories from cooking athletics genetics green chemistry and more Tests Richard C. Sweetland, 1991 Revises the information in the second edition and presents over 700 new or revised tests The Psychology section contains 20 subsections Education has 54 subsections and Business has 13 subsections Does not contain reliability validity and normative data Use the complementary *Test Critiques* series for this information Basic Research Resumés, 1961 **Chemistry** Lemay, Prentice-Hall Staff, 1996 **Descriptive Inorganic Chemistry** Geoff Rayner-Canham, Tina Overton, 2014-09-19 This bestselling text gives students a less rigorous less mathematical way of learning inorganic chemistry using the periodic table as a context for exploring chemical properties and uncovering relationships between elements in different groups The authors help students understand the relevance of the subject to their lives by covering both the historical development and fascinating contemporary applications of inorganic chemistry especially in regard to industrial processes and environmental issues The new edition offers new study tools expanded coverage of biological applications and new help with problem solving **Aeronautical Engineering Review**, 1955 Chemical Demonstrations Bassam Z. Shakhashiri, 1983 Describes and gives instructions for lecture demonstrations covering acids and bases and liquids solutions and colloids **Chemists' Guide to Effective Teaching** Norbert J. Pienta, Melanie M. Cooper, Thomas J. Greenbowe, 2005 For courses in Methods of Teaching Chemistry Useful for new professors chemical educators or students learning to teach chemistry Intended for anyone who teaches chemistry or is learning to teach it this book examines applications of learning theories presenting actual techniques and practices that respected professors have used to implement and achieve their goals Each chapter is written by a chemist who has expertise in the area and who has experience in applying those ideas in their classrooms This book is a part of the Prentice Hall Series in Educational Innovation for Chemistry **Investigating Visuospatial and Chemistry Skills Using Physical and Computer Models** Mithra Beikmohamadi, 2006 **Chemical Modeling** Alan Hinchliffe, 1999-10-18 *Chemical Modeling*

From Atoms to Liquids examines materials in terms of the basic properties of atoms molecules and polymer chains In particular the interactions between these fundamental building blocks and the intermolecular and intramolecular potentials are examined Fundamental theories of the constituent particles are covered introducing and developing classical mechanics and quantum mechanics from basics These theories are then applied to modeling developing models from both classical mechanics and quantum mechanics The book aims to make this subject both appealing and relevant whilst avoiding excessive mathematical rigour Carefully structured the text starts by introducing classical quantum and statistical mechanics before moving on to cover the modeling of solids gases and liquids To bring the subject alive many real life examples and applications have been included This book brings together the often scattered and diffuse background information that is essential for a full understanding of chemical modeling

FEATURES Assumes no prior knowledge of modeling Brings together all the background information that a reader needs to know to fully understand modeling Develops classical and quantum mechanical theories from basics Avoids unnecessary mathematical rigour Includes a detailed mathematical Toolbox as a ready reference Includes brief descriptions and web addresses of key software packages

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