

# Oxidation Numbers Worksheet

Directions: Use the *Rules for Assigning Oxidation Numbers* to determine the oxidation number assigned to each element in each of the given chemical formulas.

Rules for Assigning Oxidation Numbers	Example:
1. The oxidation number of any uncombined element is <b>0</b> .	Ne, O <sub>2</sub> = <b>0</b>
2. The oxidation number of a monatomic ion equals the charge on the ion.	Fe <sup>+3</sup> = <b>+3</b>
3. The oxidation number of halogens in a compound is mostly <b>-1</b> .	F <sup>-1</sup> = <b>-1</b>
4. Oxygen has an oxidation number of <b>-2</b> unless it's a peroxide is <b>-1</b>	H <sub>2</sub> O <sup>-2</sup> = <b>-2</b> Na <sub>2</sub> O <sub>2</sub> <sup>-2</sup> = <b>-1</b>
5. The oxidation number of a metal is <b>+1</b> in Group 1 and <b>+2</b> in Group 2.	K <sup>+1</sup> = <b>+1</b>
6. Hydrogen works with <b>+1</b> with nonmetals and <b>-1</b> with metals.	H <sup>+1</sup> Cl <sup>-1</sup> = <b>+1</b> Na <sup>+1</sup> H <sup>-1</sup> = <b>-1</b>
7. The sum of the oxidation numbers of all atoms in a neutral compound is 0.	
8. The sum of the oxidation numbers of all atoms in a polyatomic ion = the charge of the ion.	

## 1. Give oxidation numbers for the underlined atoms in these molecules:

- |  |     |     |                         |                                       |     |    |
|--|-----|-----|-------------------------|---------------------------------------|-----|----|
| a. <u>Cs<sub>2</sub>O</u>                          | Cs: | O:  | i. <u>N<sub>2</sub></u> | N:                                    |     |    |
| b. <u>N<sub>2</sub>O<sub>3</sub></u>               | N:  | O:  | j. <u>Kr</u>            | Kr:                                   |     |    |
| c. <u>Na<sub>4</sub>SiO<sub>4</sub></u>            | Na: | Si: | O:                      | k. <u>H<sub>2</sub>O</u>              | H:  | O: |
| d. <u>K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub></u> | K:  | Cr: | O:                      | l. <u>FeO</u>                         | Fe: | O: |
| e. <u>H<sub>2</sub>O<sub>2</sub></u>               | H:  | O:  | (This is a peroxide)    | m. <u>CaS</u>                         | Ca: | S: |
| f. <u>Al(OH)<sub>3</sub></u>                       | Al: | O:  | H:                      | n. <u>H<sub>2</sub></u>               | H:  |    |
| g. <u>HPO<sub>3</sub></u>                          | H:  | P:  | O:                      | o. <u>He</u>                          | He: |    |
| h. <u>H<sub>2</sub>SeO<sub>3</sub></u>             | H:  | Se: | O:                      | p. <u>H<sub>2</sub>SO<sub>4</sub></u> | S:  | H: |

## 2. Give the oxidation numbers for the following ions

- |                                       |     |                           |                             |                           |     |                           |    |    |
|---------------------------------------|-----|---------------------------|-----------------------------|---------------------------|-----|---------------------------|----|----|
| a. <u>Cu<sup>+1</sup></u>             | Cu: | b. <u>Co<sup>2+</sup></u> | Co:                         | c. <u>Cl<sup>-1</sup></u> | Cl: |                           |    |    |
| d. <u>IO<sub>2</sub><sup>-1</sup></u> | I:  | O:                        | e. <u>SbFe<sup>-1</sup></u> | Sb:                       | Fe: | f. <u>OH<sup>-1</sup></u> | O: | H: |

# Formulas And Oxidation Numbers Lab Answers

**Jo A. Beran**



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