

# CHEMISTRY MULTIPLE CHOICE QUESTIONS

## A. Atoms, molecules and stoichiometry

2002 -2019



C

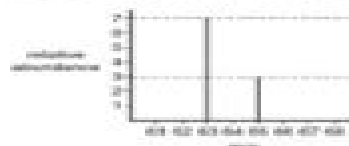
As a consequence, an adult human can be considered to have a daily diet of 1.80 kg of carbohydrate (empirical formula  $C_6H_{10}O_5$ ).

Which mass of carbon dioxide does a person produce each day if all the carbohydrate eaten is digested and oxidised?

- A 0.907 kg    B 0.906 kg    C 1.70 kg    D 3.54 kg

(2002 NA/12)

The diagram shows the mass spectrum of a sample of naturally-occurring copper.



What is the relative atomic mass of this copper?

- A 63.2    B 63.5    C 63.6    D 64.0

(2002 NA/12)

Which pairs of compounds have the same empirical formula?

- ethane and ethene
- ethane and cyclohexane
- cyclohexane and oct-1-ene

(2002 NA/12)

A stream of  $10\text{ cm}^3$  of methane and  $10\text{ cm}^3$  of ethane was sparked with an excess of oxygen. After cooling to room temperature, the residual gas was passed through aqueous potassium hydroxide.

Which compounds have the empirical formula  $CH_2O$ ?

- methanol
- ethanoic acid
- methyl methanoate

(2002 ON/11)

What is the number of molecules in  $500\text{ cm}^3$  of oxygen under room conditions?

- A  $1.25 \times 10^{23}$   
 B  $1.25 \times 10^{24}$   
 C  $3.0 \times 10^{23}$   
 D  $3.0 \times 10^{24}$

(2004 MO/12)

In the preparation of soft margarine, glycerol trioleate



is fully hydrogenated so that, on average, one of its side-chains is converted into the  $\text{CH}_3(\text{CH}_2)_7\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$  moiety and two side-chains are converted into the  $\text{CH}_3(\text{CH}_2)_7\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$  moiety.

How many moles of hydrogen are required to convert one mole of glycerol trioleate into the soft margarine?

- A 4    B 5    C 6    D 8

(2008 MO/12)

Use of the Data Booklet is relevant to this question.

Analysed classes may obtain more credit awards of entry levels. Award is a 100% and

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**S Ashworth**



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