

Molarity = Moles of solute / Liters of Solution (abbreviation = M)

Molality = Moles of solute / Kg of Solvent (abbreviation = m)

Normality = number of equivalent of solute x Molarity of Solution (abbreviation = N)

Mass Percent = mass of solute / mass of solution

1. How many moles of ethyl alcohol,  $C_2H_5OH$ , are present in 65 mL of a 1.5 M solution?

$$\frac{65 \text{ mL } C_2H_5OH}{1000 \text{ mL } C_2H_5OH} \times \frac{1 \text{ L } C_2H_5OH}{1 \text{ L } C_2H_5OH} \times \frac{1.5 \text{ mole } C_2H_5OH}{1 \text{ L } C_2H_5OH} = 0.098 \text{ moles } C_2H_5OH$$

2. How many liters of a 6.0 M solution of acetic acid  $CH_3COOH$ , contain 0.0030 moles of acetic acid?

$$\frac{0.0030 \text{ moles } CH_3COOH}{6.0 \text{ moles } CH_3COOH} \times \frac{1 \text{ L solution}}{1 \text{ L solution}} = 5.0 \times 10^{-4} \text{ L solution}$$

3. You want 85 g of KOH. How much of a 3.0 m solution of KOH will provide it?

$$\frac{85 \text{ g KOH}}{56.11 \text{ g KOH}} \times \frac{1 \text{ mole}}{1 \text{ mole}} \times \frac{1 \text{ kg solvent}}{3.0 \text{ mole KOH}} \times \frac{1000 \text{ g solvent}}{1 \text{ kg solvent}} = 5.0 \times 10^2 \text{ g solvent}$$

$$5.0 \times 10^2 \text{ g solvent} + 85 \text{ g KOH} = 590 \text{ g solution}$$

4. If you dissolve 0.70 moles of HCl in enough water to prepare 250 mL of solution, what is the molarity of the solution you have prepared?

$$M = \frac{0.70 \text{ moles HCl}}{0.250 \text{ L solution}} = 2.8 \text{ M HCl}$$

5. A solution is prepared by adding 2.0 L of 6.0 M HCl to 500 mL of a 9.0 M HCl solution. What is the molarity of the new solution? (Remember, the volumes are additive)

$$\frac{2.0 \text{ L solution}}{1 \text{ L solution}} \times \frac{6.0 \text{ moles HCl}}{1 \text{ L solution}} = 12 \text{ moles HCl}$$

$$\frac{0.500 \text{ L solution}}{1 \text{ L solution}} \times \frac{9.0 \text{ moles HCl}}{1 \text{ L solution}} = 4.5 \text{ moles HCl}$$

$$12 \text{ moles HCl} + 4.5 \text{ moles HCl} = 17 \text{ moles}$$

$$2.0 \text{ L solution} + 0.500 \text{ L solution} = 2.5 \text{ L solution}$$

$$M = \frac{17 \text{ moles}}{2.5 \text{ L solution}} = 6.8 \text{ M HCl}$$

6. Convert the following Molarities to Normalities.

a. 2.5 M HCl = 2.5 N

b. 1.4 M  $H_2SO_4$  = 2.8 N

c. 1.0 M NaOH = 1.0 N

d. 0.5 M  $Ca(OH)_2$  = 1 N

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## **Table of Contents Molality Answers Key**

1. Understanding the eBook Molality Answers Key
  - The Rise of Digital Reading Molality Answers Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Molality Answers Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Molality Answers Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Molality Answers Key
  - Personalized Recommendations
  - Molality Answers Key User Reviews and Ratings
  - Molality Answers Key and Bestseller Lists
5. Accessing Molality Answers Key Free and Paid eBooks
  - Molality Answers Key Public Domain eBooks
  - Molality Answers Key eBook Subscription Services

- Molality Answers Key Budget-Friendly Options
- 6. Navigating Molality Answers Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Molality Answers Key Compatibility with Devices
  - Molality Answers Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Molality Answers Key
  - Highlighting and Note-Taking Molality Answers Key
  - Interactive Elements Molality Answers Key
- 8. Staying Engaged with Molality Answers Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Molality Answers Key
- 9. Balancing eBooks and Physical Books Molality Answers Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Molality Answers Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Molality Answers Key
  - Setting Reading Goals Molality Answers Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Molality Answers Key
  - Fact-Checking eBook Content of Molality Answers Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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