

7.3

BONDING IN METALS

Section Review

Objectives

- 1. Identify metallic bonding characteristics
- 2. Describe the delocalized electron model
- 3. Explain the properties of metals

Vocabulary

- metallic bond
- metal

Part A Completion

Use the vocabulary words in parentheses to complete each sentence. Write the word or phrase in the space provided.

1. The _____ of metals is due to the _____ of valence electrons. (delocalized electron model)
2. _____ is the property of metals that allows them to be drawn into wires. (ductility)
3. _____ is the property of metals that allows them to be hammered into thin sheets. (malleability)
4. _____ is the property of metals that allows them to be drawn into thin sheets. (ductility)
5. _____ is the property of metals that allows them to be drawn into thin sheets. (ductility)
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19. _____ is the property of metals that allows them to be drawn into thin sheets. (ductility)
20. _____ is the property of metals that allows them to be drawn into thin sheets. (ductility)

Part B True/False

Write *true* or *false* in the space provided for each statement.

1. _____ Metals are good conductors of electricity. (true)
2. _____ Metals are good conductors of heat. (true)
3. _____ Metals are good conductors of heat. (true)
4. _____ Metals are good conductors of heat. (true)
5. _____ Metals are good conductors of heat. (true)
6. _____ Metals are good conductors of heat. (true)
7. _____ Metals are good conductors of heat. (true)
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18. _____ Metals are good conductors of heat. (true)
19. _____ Metals are good conductors of heat. (true)
20. _____ Metals are good conductors of heat. (true)

Bonding In Metals Section Review Answers Key

S Ashworth



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